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### Correlation Between Infrared Intensities, Dissociation Energies, and Equilibrium Internuclear Distances for Diatomic Molecules\*

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ONE may expand the dipole moment as a function of inter-nuclear distance as follows:

$$\mu = \mu_0 + \left(\frac{\partial\mu}{\partial r}\right)_{r=r_0}(r-r_0) + \dots \quad (1)$$

The derivative  $(\partial\mu/\partial r)_{r=r_0}$  may be obtained from infrared intensity measurements. Using the best available data of  $(\partial\mu/\partial r)_{r=r_0}$  for NO, CO, HCl, and HBr,<sup>1</sup> and the most generally accepted values for the dissociation energies and equilibrium internuclear distances of these diatomic molecules,<sup>2</sup>  $(\partial\mu/\partial r)_{r=r_0}$  is found to be proportional to the dissociation energy divided by the square of the equilibrium internuclear distance. A graph of  $(\partial\mu/\partial r)_{r=r_0}$  as a function of  $E/r_0^2$  is given in Fig. 1. The values used for  $E$ ,  $r_0$ , and  $(\partial\mu/\partial r)_{r=r_0}$  are given in Table I. It is pleasant to observe

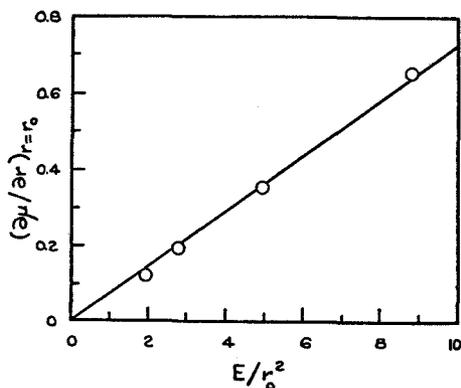


FIG. 1. Graph of  $(\partial\mu/\partial r)_{r=r_0}$  in e (electronic charge units) as a function of  $E/r_0^2$  (ev/A<sup>2</sup>).

TABLE I. Molecular parameters.

	$\left(\frac{\partial\mu}{\partial r}\right)_{r=r_0}$	$r_0$ (A)	$E$ (ev)
HBr	0.121e	1.414	3.75
HCl	0.198e	1.275	4.43
NO	0.352e	1.151	6.48
CO	0.655e	1.128	11.11

that the data may be correlated so that a straight line goes through the origin, which, in a sense, constitutes a fifth plotted point.

The higher values of CO and NO, which are perhaps more widely accepted, were used here. It is interesting that the value of  $(\partial\mu/\partial r)_{r=r_0}$  reported by Penner and Weber for NO, which differs appreciably from previously reported values, correlates with these results. It is further interesting to note that the constant of proportionality has a value of approximately one electronic charge per cubic angstrom.

A fair correlation for overtone intensities on HBr, HCl, and CO has also been noted by plotting one of the possible values of  $(\partial^2\mu/\partial r^2)_{r=r_0}$  against  $E/r_0^3$ .<sup>3</sup> These relations should be useful for estimating infrared intensities of molecules such as HF and OH, where measurement is difficult.

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<sup>1</sup> S. S. Penner and D. Weber, J. Chem. Phys. 21, 649 (1953); 19, 1351 (1951).

<sup>2</sup> G. Herzberg, *Spectra of Diatomic Molecules* (D. Van Nostrand Company, Inc., New York, 1950); A. Gaydon, *Dissociation Energies and Spectra of Diatomic Molecules* (John Wiley and Sons, Inc., New York, 1947).

<sup>3</sup> In the paper of S. S. Penner and D. Weber, J. Chem. Phys. 21, 649 (1953), the coefficients of  $\xi^2$  in Eq. (4) for NO should be changed to  $8.77 \times 10^{-18}$  or  $-3.06 \times 10^{-18}$ ; in Eq. (5) for HCl to  $2.96 \times 10^{-18}$  or  $-0.140 \times 10^{-18}$ . The indicated corrections are the result of recalculation of appropriate coefficients using (a) the intensity data given in the original paper and (b), other required spectroscopic constants listed by Herzberg, reference 2.