

# **Supplementary Information: Sulfur monoxide dimer chemistry as a possible source of polysulfur in the upper atmosphere of Venus**

**Joseph P. Pinto<sup>1</sup>, Jiazheng Li<sup>2</sup>, Franklin P. Mills<sup>3,4</sup>, Emmanuel Marcq<sup>5</sup>, Daria Evdokimova<sup>5,6</sup>, Denis Belyaev<sup>6</sup> and Yuk L. Yung<sup>2,7</sup>**

<sup>1</sup>University of North Carolina at Chapel Hill, NC, USA

<sup>2</sup>Division of Geological and Planetary Science, California Institute of Technology, Pasadena, CA, USA

<sup>3</sup>Australian National University, Canberra, ACT, Australia

<sup>4</sup>Space Science Institute, Boulder, CO, USA

<sup>5</sup>LATMOS/CNRS/Sorbonne Université/UVSQ, France

<sup>6</sup>Space Research Institute of the Russian Academy of Sciences (IKI), Moscow, Russia

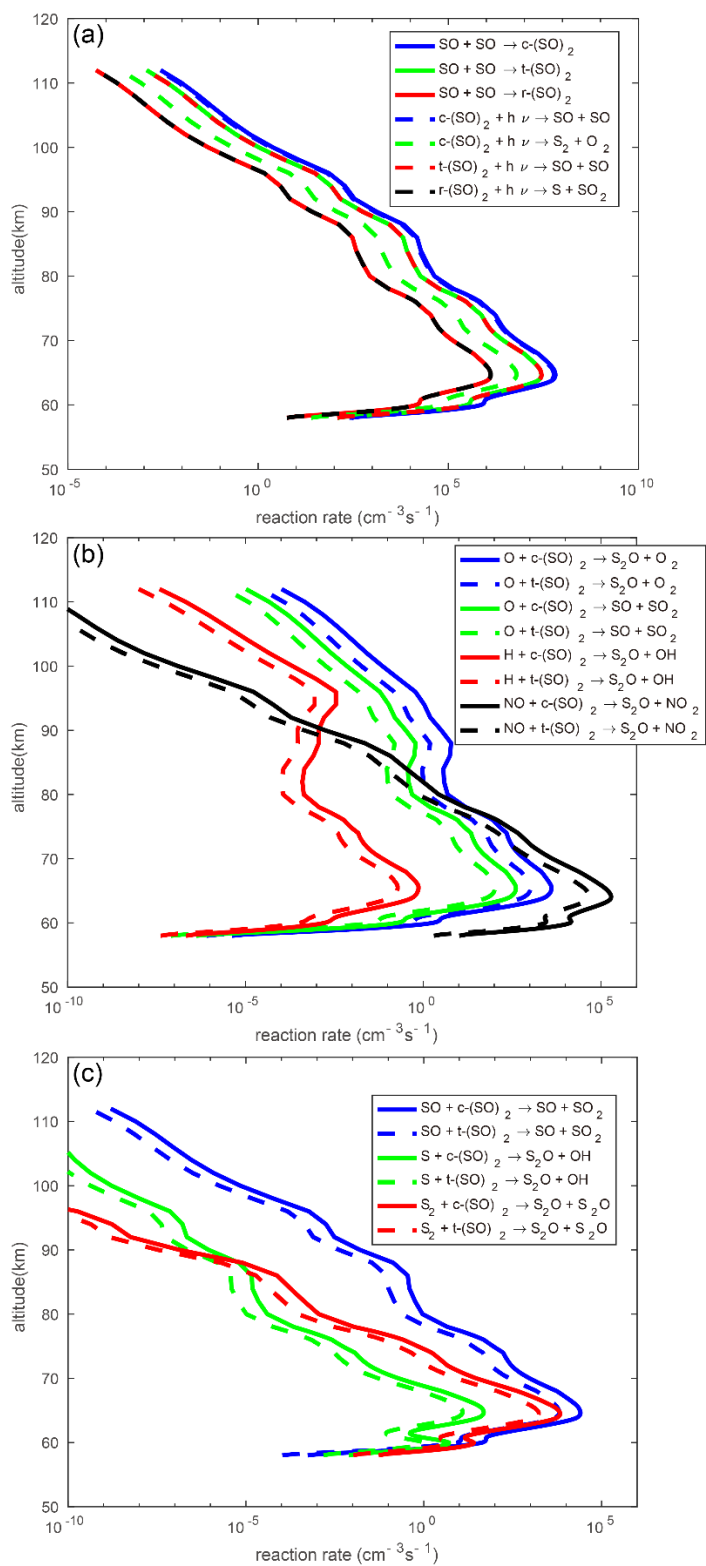
<sup>7</sup>Jet Propulsion Laboratory, California Institute of Technology, Pasadena, CA, USA

Corresponding author: Jiazheng Li (jiazheng@caltech.edu)

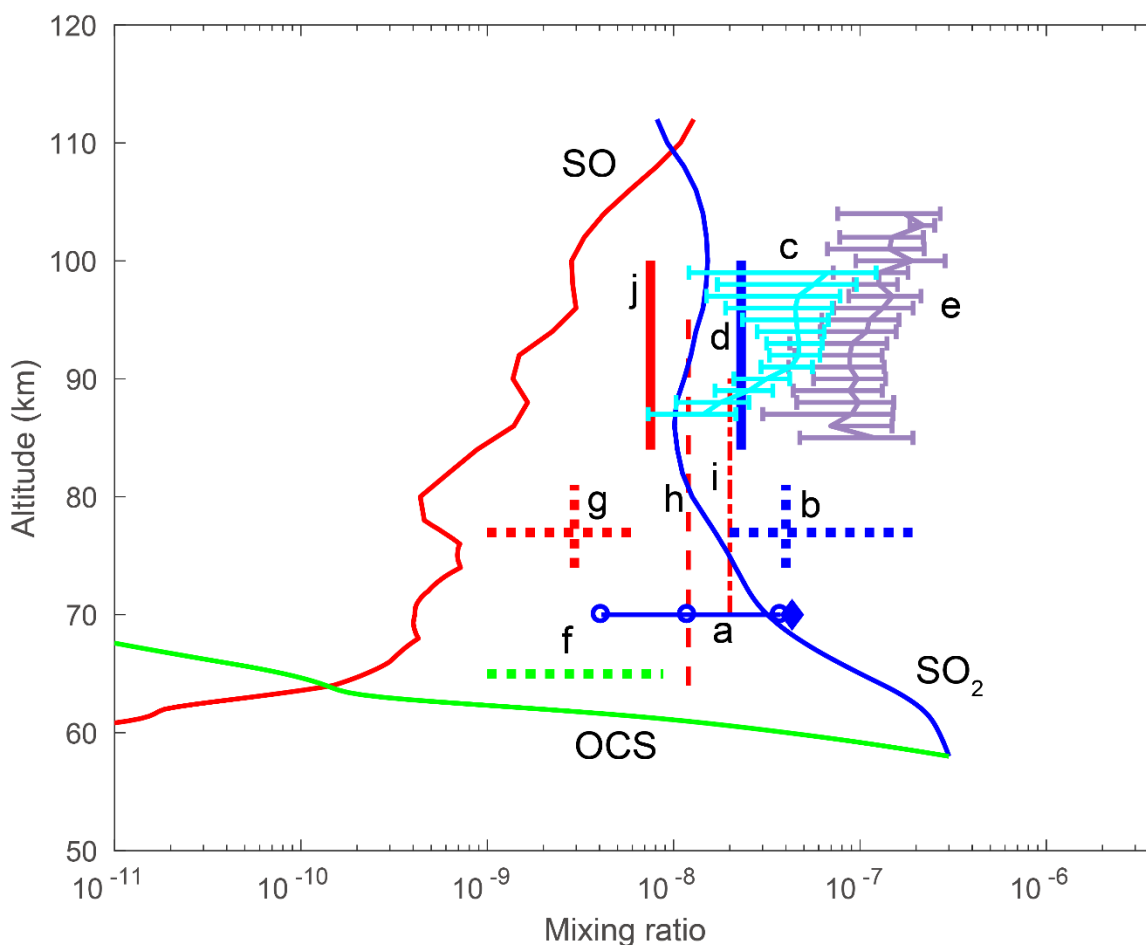
**Supplementary Table 1.** Reactions involving the SO dimer in the photochemical model.

Reactions	Photolysis Rate, J Rate Coefficients, k*	References
1. $\text{SO} + \text{SO} + \text{M} \rightarrow \text{c}-(\text{SO})_2 + \text{M}$	$k = 2.15 \times 10^{-31}; 1.00 \times 10^{-11}$	a
2. $\text{SO} + \text{SO} + \text{M} \rightarrow \text{t}-(\text{SO})_2 + \text{M}$	$k = 2.15 \times 10^{-31}; 1.00 \times 10^{-11}$	a
3. $\text{SO} + \text{SO} + \text{M} \rightarrow \text{r}-(\text{SO})_2 + \text{M}$	$k = 8.80 \times 10^{-33}; 1.00 \times 10^{-11}$	a
4. $\text{c}-(\text{SO})_2 + h\nu \rightarrow \text{SO} + \text{SO}$	$J = 9.30 \times 10^{-2}$	a
5. $\text{c}-(\text{SO})_2 + h\nu \rightarrow \text{S}_2 + \text{O}_2$	$J = 9.30 \times 10^{-3}$	a
6. $\text{t}-(\text{SO})_2 + h\nu \rightarrow \text{SO} + \text{SO}$	$J = 9.30 \times 10^{-2}$	a
7. $\text{r}-(\text{SO})_2 + h\nu \rightarrow \text{S} + \text{SO}_2$	$J = 1.20 \times 10^{-1}$	a
8. $\text{O} + \text{c}-(\text{SO})_2 \rightarrow \text{S}_2\text{O} + \text{O}_2$	$k = 3.00 \times 10^{-14}$	a
9. $\text{O} + \text{c}-(\text{SO})_2 \rightarrow \text{SO} + \text{SO}_2$	$k = 3.00 \times 10^{-15}$	a
10. $\text{O} + \text{t}-(\text{SO})_2 \rightarrow \text{S}_2\text{O} + \text{O}_2$	$k = 3.00 \times 10^{-14}$	a
11. $\text{O} + \text{t}-(\text{SO})_2 \rightarrow \text{SO} + \text{SO}_2$	$k = 3.00 \times 10^{-15}$	a
12. $\text{H} + \text{c}-(\text{SO})_2 \rightarrow \text{S}_2\text{O} + \text{OH}$	$k = 3.00 \times 10^{-14}$	b
13. $\text{H} + \text{t}-(\text{SO})_2 \rightarrow \text{S}_2\text{O} + \text{OH}$	$k = 3.00 \times 10^{-14}$	b
14. $\text{NO} + \text{c}-(\text{SO})_2 \rightarrow \text{S}_2\text{O} + \text{NO}_2$	$k = 3.00 \times 10^{-14}$	b
15. $\text{NO} + \text{t}-(\text{SO})_2 \rightarrow \text{S}_2\text{O} + \text{NO}_2$	$k = 3.00 \times 10^{-14}$	b
16. $\text{SO} + \text{c}-(\text{SO})_2 \rightarrow \text{S}_2\text{O} + \text{SO}_2$	$k = 3.00 \times 10^{-14}$	b
17. $\text{SO} + \text{t}-(\text{SO})_2 \rightarrow \text{S}_2\text{O} + \text{SO}_2$	$k = 3.00 \times 10^{-14}$	a
18. $\text{S} + \text{c}-(\text{SO})_2 \rightarrow \text{S}_2\text{O} + \text{SO}$	$k = 3.00 \times 10^{-14}$	a
19. $\text{S} + \text{t}-(\text{SO})_2 \rightarrow \text{S}_2\text{O} + \text{SO}$	$k = 3.00 \times 10^{-14}$	a
20. $\text{S}_2 + \text{c}-(\text{SO})_2 \rightarrow 2\text{S}_2\text{O}$	$k = 3.00 \times 10^{-14}$	a
21. $\text{S}_2 + \text{t}-(\text{SO})_2 \rightarrow 2\text{S}_2\text{O}$	$k = 3.00 \times 10^{-14}$	a

\*Low and high pressure rate coefficients given for SO dimer formation; J values given at 112 km, units are  $\text{cm}^{-3}\text{s}^{-1}$  for J and  $\text{cm}^3\text{s}^{-1}\text{molec}^{-1}$ ,  $\text{cm}^6\text{s}^{-1}\text{molec}^{-2}$  for k. References: (a) ref. 1 and references therein; (b) estimated based on ref. 1.



**Supplementary Figure 1.** Reaction rate profiles of the reactions in Table S1.



**Supplementary Figure 2.** Comparison of modeled profiles of  $\text{SO}_2$ , OCS and SO with observations for  $\text{SO}_2 = 0.3$  ppm and OCS = 0.3 ppm at 58 km, the lower boundary of the model. Model profiles are shown as thin solid lines; (blue)  $\text{SO}_2$ , (green) OCS, (red) SO. Model  $\text{SO}_2$  and SO profiles considering new  $\text{H}_2\text{SO}_4$  cross sections from ref. 2 are shown as thin dotted lines. Model OCS profile using 100 ppm at 58 km shown as dotted line. Data sources: (a)  $\text{SO}_2$ , interquartile range from 2010 through 2014, diamond shows mean of distribution, ref. 3; (b)  $\text{SO}_2$ , ref. 4; (c)  $\text{SO}_2$ , ref. 5 solar occultation with 1- $\sigma$  error bars; (d)  $\text{SO}_2$ , ref. 6; (e)  $\text{SO}_2$ , ref. 5 stellar occultation with 1- $\sigma$  error bars; (f) OCS, ref. 7; (g) SO, ref. 4; (h) SO, ref. 8; (i) SO, ref. 9; (j) SO, ref. 6.

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