

SUPPLEMENTARY INFORMATION

# Overcoming Hurdles in Oxygen Evolution Catalyst Discovery via Co-Design

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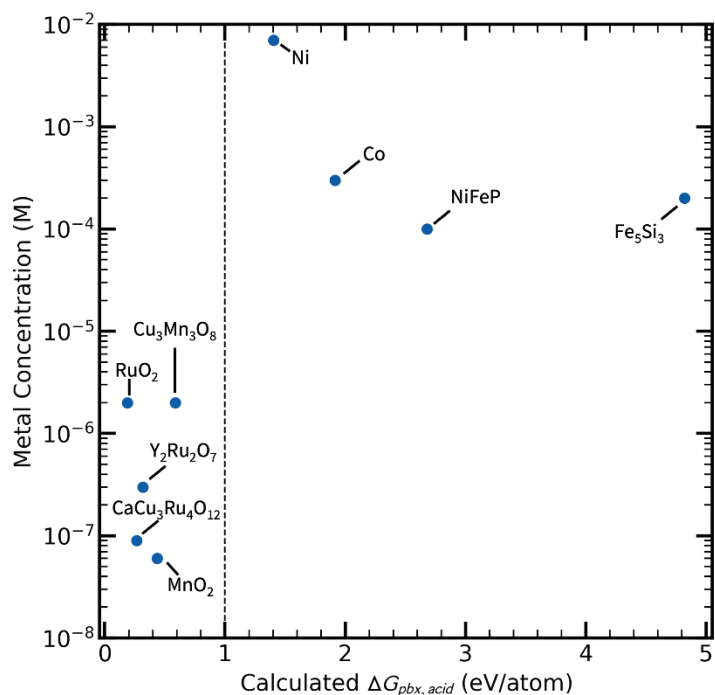


Figure S1: Establishment of a reasonable upper limit of  $\Delta G_{pbx}$  for operationally-stable catalysts for OER in acid. Using catalysts from Figure 4, the experimentally measured dissolved metal concentrations are shown as a function of the calculated  $\Delta G_{pbx,acid}$  (pH=1). Only materials with a known computational bulk structure are shown. Metal concentrations below  $10^{-5}$  M are observed only for materials below the  $\Delta G_{pbx}$  threshold of 1 eV/metal atom (vertical dashed line).

Table S1: List of  $d_0$  descriptors for each of the transition metal elements considered in this study.

Element	Acid (pH= 1)		Alkaline (pH=14)	
	$d_0$	End Member	$d_0$	End Member
Sc	0	$\text{Sc}^{+3}_{(\text{aq})}$	0	$\text{Sc}_2\text{O}_3$
Ti	0	$\text{TiO}_2$	0	$\text{TiO}_2$
V	0	$\text{VO}_4^{-}_{(\text{aq})}$	0	$\text{VO}_4^{-}_{(\text{aq})}$
Cr	0	$\text{CrHO}_4^{-}$	0	$\text{CrO}_4^{2-}$
Mn	3	$\text{MnO}_2$	0	$\text{MnO}_4^{-}$
Fe	4	$\text{Fe}^{+3}_{(\text{aq})}$	4	$\text{Fe}_2\text{O}_3$
Co	7	$\text{Co}^{+2}_{(\text{aq})}$	6	$\text{CoOOH}$ , $\text{Li}(\text{CoO}_2)_2$ , $\text{LaCoO}_3$
Ni	8	$\text{Ni}^{+2}_{(\text{aq})}$	8	$\text{Ni}(\text{OH})_3^{-}_{(\text{aq})}$ , $\text{LaNiO}_3$
Cu	9	$\text{Cu}^{+2}_{(\text{aq})}$	8	$\text{Cu}_2\text{O}_3$ , $\text{LaCuO}_3$ , $\text{Zn}(\text{CuO}_2)_2$ , $\text{Ca}(\text{CuO}_2)_2$
Zn	10	$\text{Zn}^{+2}_{(\text{aq})}$	10	$\text{ZnO}_2^{-2}_{(\text{aq})}$ , $\text{Zn}(\text{CuO}_2)_2$
Y	0	$\text{Y}^{+3}_{(\text{aq})}$	0	$\text{Y}_2\text{O}_3$
Zr	0	$\text{ZrO}^{+2}_{(\text{aq})}$	0	$\text{ZrO}_2$
Nb	0	$\text{Nb}(\text{HO})_5_{(\text{aq})}$	0	$\text{NbO}_3^{-}_{(\text{aq})}$
Mo	0	$\text{MoO}_3$	0	$\text{MoO}_4^{-2}_{(\text{aq})}$
Tc	0	$\text{Tc}_2\text{O}_7$	0	$\text{TcO}_4^{-}_{(\text{aq})}$
Ru	0	$\text{RuO}_4$	0	$\text{RuO}_4$
Rh	5	$\text{RhO}_2$	5	$\text{RhO}_2$
Pd	6	$\text{PdO}_2$	6	$\text{PdO}_2$
Ag	10	$\text{Ag}^{+}_{(\text{aq})}$	9	$\text{AgO}$
Cd	10	$\text{Cd}^{+2}_{(\text{aq})}$	10	$\text{Cd}(\text{OH})_2_{(\text{aq})}$
Hf	0	$\text{Hf}^{+4}_{(\text{aq})}$	0	$\text{HfO}_2$
Ta	2	$\text{TaO}_2^{+}_{(\text{aq})}$	0	$\text{Ta}_2\text{O}_5$ , $\text{SrTa}_2\text{O}_6$

W	0	WO <sub>3</sub>	0	WO <sub>4</sub> <sup>-2</sup> <sub>(aq)</sub>
Re	0	ReO <sub>4</sub> <sup>-</sup> <sub>(aq)</sub>	0	ReO <sub>4</sub> <sup>-</sup> <sub>(aq)</sub>
Os	0	OsO <sub>4</sub>	0	OsO <sub>4</sub>
Ir	2	IrO <sub>4</sub> <sup>-</sup> <sub>(aq)</sub>	2	IrO <sub>4</sub> <sup>-</sup> <sub>(aq)</sub>
Pt	6	PtO <sub>2</sub> / BiPt <sub>3</sub> O <sub>11</sub>	6	BiPt <sub>3</sub> O <sub>11</sub> , La <sub>3</sub> Pt <sub>3</sub> O <sub>11</sub> , PtO <sub>2</sub>
Au	9	H <sub>3</sub> AuO <sub>3</sub> <sub>(aq)</sub>	9	H <sub>2</sub> AuO <sub>3</sub> <sup>-</sup> <sub>(aq)</sub> , Bi <sub>2</sub> AuO <sub>5</sub> , Sr(AuO <sub>2</sub> ) <sub>2</sub> , LaAuO <sub>3</sub>

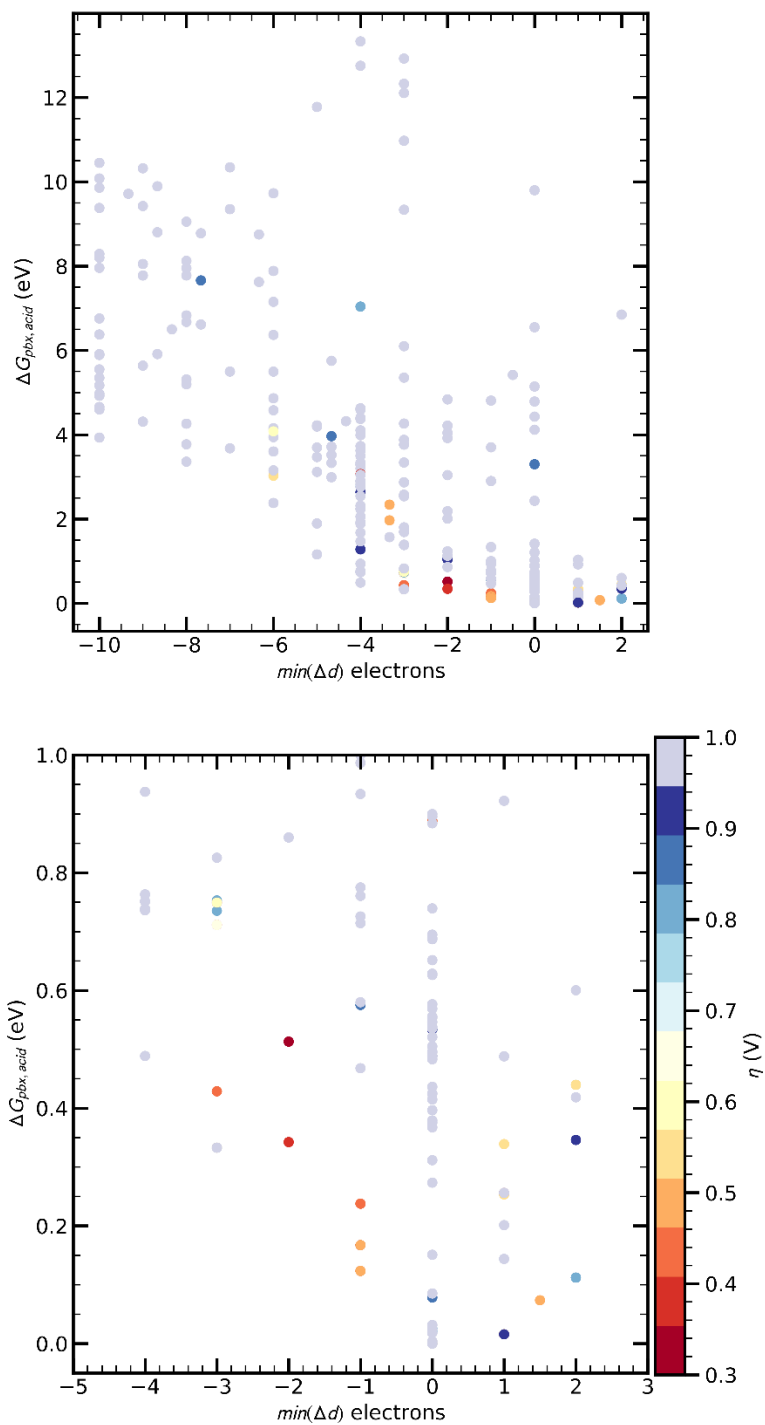


Figure S2: The analogue of Figure 6 using an alternative acid stability descriptor, the minimum  $\Delta d$  among all cations in the catalyst. The bottom panel has a maximum  $\Delta G_{pbx}$  of 1 eV/atom, for which all considered materials have  $\min(\Delta d)$  no lower than -4 and no greater than 2.