

CHEMISTRY OF CHLORINE IN DENSE INTERSTELLAR CLOUDS

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ABSTRACT

Laboratory experiments and theoretical modeling show that the chemistry of chlorine is fairly simple in dense interstellar clouds, with Cl and HCl as the only species whose fractional abundances are significant. The estimated fraction of gas-phase chlorine present as HCl lies between 25%–65%, in good agreement with the recent observations of the ground state HCl transition by Blake, Keene, and Phillips. These results, combined with the observational limits on HCl, indicate that chlorine is not severely depleted in dense interstellar clouds.

Subject headings: interstellar: abundances — interstellar: molecules

I. INTRODUCTION

Elemental chlorine has been well studied in diffuse clouds by the *Copernicus* satellite, using optical and UV transitions of Cl I and Cl II, which have shown that chlorine is depleted by no more than a factor of 3 below the solar abundance ($\sim 3 \times 10^{-7}$) in such clouds (Jura and York 1978). Searches for HCl and other molecular species have produced only upper limits to date (Smith *et al.* 1980), in contrast to some calculations (Jura 1974; Black and Dalgarno 1977) which suggested that detectable column densities of HCl should exist. More recent theoretical models now reproduce these observations quite well and indicate that molecular chlorine compounds do not play a significant role in diffuse clouds (van Dishoeck 1984). The chemistry of chlorine in dense interstellar clouds remains rather more uncertain, however. As one of the more volatile second-row elements, the gas-phase abundance of chlorine in these predominantly molecular clouds is expected to remain high. The ubiquitous nature of H₂ in dense clouds led early model calculations to conclude that HCl would be the dominant, or sole, gas-phase chlorine reservoir in such objects and that rotational line emission from hydrogen chloride would contribute substantially to the energy balance of molecular clouds (Dalgarno *et al.* 1974). The large rotational constants of HCl place its pure rotational spectrum in the submillimeter and far-infrared wavelength regions in which high-resolution astronomical observations are both difficult and sparse.

Recently, however, Blake, Keene, and Phillips (1985) have successfully detected the HCl ground-state transition at 625.9 GHz toward the Orion molecular cloud. These observations indicate that while the amount of gas-phase hydrogen chloride is large, it is somewhat below that expected for an undepleted chlorine abundance fully associated into HCl. In order to determine whether depletion mechanisms alone are responsible for this decrease, or whether chemical processes also limit the HCl abundance, we have carried out laboratory experiments and theoretical model calculations on the ion-molecule chemistry of interstellar chlorine. These studies have shown that HCl is a major, but not the sole, reservoir of

gaseous chlorine in dense interstellar clouds, with neutral atomic chlorine remaining as the other significant species. We also find that, although higher than in diffuse clouds, the chlorine depletion is not severe in dense molecular clouds.

II. CHEMICAL DATA

Many of the reactions of importance to interstellar chlorine chemistry have not been previously measured in the laboratory. We have therefore measured a number of these reactions, particularly those involving CCl⁺, using the ion cyclotron resonance techniques discussed in Thorne, Anicich, and Huntress (1983). Table 1 summarizes the laboratory work

TABLE 1
LABORATORY RATE CONSTANTS FOR ION-MOLECULE
REACTIONS INVOLVING CHLORINE

Reactions	Rate Constant <i>k</i> (10 ⁻⁹ cm ³ s ⁻¹)
Reactions of H ₂ Cl ⁺ :	
H ₂ Cl ⁺ + CO → HCO ⁺ + HCl.....	0.42
H ₂ Cl ⁺ + H ₂ O → H ₃ O ⁺ + HCl.....	2.00
Reactions of CCl ⁺ :	
CCl ⁺ + H ₂ → no reaction.....	≤0.01
CCl ⁺ + N ₂ → no reaction.....	≤0.03
CCl ⁺ + O ₂ → no reaction.....	≤0.03
CCl ⁺ + CO → no reaction.....	≤0.03
CCl ⁺ + CO ₂ → no reaction.....	≤0.03
CCl ⁺ + H ₂ O → no reaction.....	≤0.03
CCl ⁺ + NH ₃ → HCNH ⁺ + HCl.....	1.25
CCl ⁺ + NH ₃ → H ₃ CN ⁺ + Cl.....	0.05
CCl ⁺ + CH ₄ → no reaction.....	≤0.03
CCl ⁺ + HCN → no reaction.....	≥0.03
CCl ⁺ + H ₂ CO → H ₂ CCl ⁺ + CO.....	0.53
Reactions of H ₂ CCl ⁺ :	
H ₂ CCl ⁺ + H ₂ → no reaction.....	≤0.01
H ₂ CCl ⁺ + CO → no reaction.....	≤0.01
Miscellaneous:	
HCl + H ₃ ⁺ → H ₂ Cl ⁺ + H ₂	3.50
HCl + HCO ⁺ → H ₂ Cl ⁺ + CO.....	0.30
HCl + CH ₃ ⁺ → H ₂ CCl ⁺ + H ₂	0.19

TABLE 2
HEATS OF FORMATION OF SPECIES CONTAINING
CHLORINE

Ions	ΔH_f^0 (kCal)	Neutrals	ΔH_f^0 (kCal)
Cl ⁺	327.7 ^a	Cl	29.1 ^a
HCl ⁺	271.8 ^a	HCl	-22.1 ^a
H ₂ Cl ⁺	~205. ^b
CCl ⁺	~340. ^c	CCl	~122. ^d
H ₂ CCl ⁺	~227. ^{a,e}

^a Rosenstock *et al.* 1977.

^b Calculated from proton affinities listed by Hartman *et al.* 1979.

^c Calculated from the heat of formation CCl⁺ and the upper limit determined for the ionization potential of CCl⁺.

^d Calculated from average or estimated bond energies.

^e Levin and Lias 1982.

performed on the ion-molecule reactions of chlorine. The CCl⁺ and H₂CCl⁺ ions were generated by electron impact on chloroform and dichloromethane, while H₂Cl⁺ was produced via the reaction of H₃⁺ and HCl. The rapid proton transfer reactions from H₂Cl⁺ and the reactions leading to and involving CCl⁺ and H₂CCl⁺ were neglected in the earliest dense cloud models (Dalgarno *et al.* 1974), but are quite important to the ion-molecule chemistry of chlorine. These reactions are the key result of the laboratory study.

The theoretical calculation used to model interstellar chlorine chemistry requires the rate coefficients of more reactions than can be currently measured, particularly those involving the reactive atomic species C and O. Thermodynamic data and chemical intuition have therefore been used in place of laboratory measurements to predict the allowed reaction pathways and most probable products of processes involving these reactants. Listed in Table 2 are the heats of formation (ΔH_f^0) used to determine the reaction energetics for those pathways included in the model which have not been measured in the laboratory. Heats of formation for the Cl⁺ and HCl⁺ ions have been calculated from the ionization potentials (IP) and ΔH_f^0 of the corresponding neutrals, while that of H₂Cl⁺ has been calculated using the measured proton affinity of HCl.

No accepted values of ΔH_f^0 exist for CCl⁺ or H₂CCl⁺. An upper limit of $\Delta H_f^0(\text{CCl}^+) \lesssim 340$ kCal per mole has been obtained by examining charge transfer reactions between CCl⁺ and a number of low IP neutrals such as NO, CH₃NH₂, and Cl₂. No charge transfer was detected between any of these species, implying an upper limit of 9.26 eV for IP(CCl), determined from NO. The low IPs of CCl and H₂CCl are consistent with the observed stability and low reactivity of the CCl⁺ and H₂CCl⁺ ions, which is explained in part by the fact that these ions are isoelectronic with the familiar and stable interstellar species CS and H₂CS.

III. THE MODEL

The procedure used to examine the interstellar chemistry of chlorine is essentially identical to that described by Thorne *et al.* (1984) in their study of interstellar phosphorus. The low fractional abundance of chlorine insures that, like phosphorus, its chemistry acts only as a small perturbation on the overall composition and chemistry of molecular clouds. Because of the low chlorine fractional abundance and the current limited

understanding of chlorine ion-molecule chemistry, we have constructed a simple kinetic model in place of a large-scale theoretical calculation. This model employs the results of previous laboratory, astronomical, and theoretical modeling efforts to effectively decouple the interstellar chlorine chemistry from those processes occurring among the more abundant first-row elements. The Runge-Kutta procedure used to examine interstellar phosphorus has been implemented on a VAX/11-780 computer to solve the set of kinetic equations listed in Table 3. Exothermic reactions included in the model which have not been measured experimentally are assigned rate coefficients of $1 \times 10^{-9} \text{ cm}^3 \text{ s}^{-1}$. These would represent upper limits and therefore introduce the maximum effect of the predicted channels.

The reactant abundances and density-dependent variables required by the kinetic procedure are listed in Table 4. Only the chlorine-containing species listed in Table 2 have their abundances calculated by the model; the fractional abundances of important neutral atoms and molecules are assumed to be density independent and are fixed to the values determined from observational data, relying mainly on the broad-band spectral-line searches of the Orion molecular cloud recently completed at the Owens Valley and Onsala Radio Observatories (Sutton *et al.* 1985; Johansson *et al.* 1984). Abundances for currently unobservable species are set to the values obtained by comprehensive chemical models of molecu-

TABLE 3
CHEMICAL EQUATIONS USED IN MODEL

Reactions	Rate Constant <i>k</i> (10 ⁻⁹ cm ³ s ⁻¹)
Cl + cosmic rays → Cl ⁺ + e ⁻	1.0 × 10 ^{-8 a}
Cl + H ⁺ ^b → Cl ⁺ + H	1.0 ^a
Cl + H ₃ ⁺ ^b → HCl ⁺ + H ₂	1.0 ^a
HCl + H ⁺ ^b → HCl ⁺ + H	4.0 ^a
HCl + He ⁺ → Cl ⁺ + H + He ...	3.3 ^c
HCl + H ₃ ⁺ → H ₂ Cl ⁺ + H ₂	3.5 ^d
HCl + HCO ⁺ → H ₂ Cl ⁺ + CO	0.0004 ^e
HCl + C ⁺ ^b → CCl ⁺ + H	1.0 ^f
HCl + CH ₃ ⁺ → H ₂ CCl ⁺ + H ₂	0.13 ^{d,e}
Cl ⁺ + H ₂ ^b → HCl ⁺ + H	1.0 ^g
HCl ⁺ + e ⁻ → Cl + H	300. ^a
HCl ⁺ + H ₂ ^b → H ₂ Cl ⁺ + H	1.30 ^g
H ₂ Cl ⁺ + e ⁻ → Cl + H ₂	150. ^a
H ₂ Cl ⁺ + e ⁻ → HCl + H	150. ^b
H ₂ Cl ⁺ + C ^b → CCl ⁺ + H ₂	0.5 ^a
H ₂ Cl ⁺ + C ^b → HCl + CH ⁺	0.5 ^a
H ₂ Cl ⁺ + S → HCl + HS ⁺	1.0 ^a
H ₂ Cl ⁺ + CO ^b → HCl + HCO ⁺	0.73 ^{d,e}
H ₂ Cl ⁺ + H ₂ O → HCl + H ₃ O ⁺	0.73 ^d
CCl ⁺ + e ⁻ → Cl + C	300. ^a
CCl ⁺ + C → Cl + C ₂ ⁺	1.0 ^a
CCl ⁺ + N → Cl + CN ⁺	1.0 ^a
CCl ⁺ + O ^b → Cl + CO ⁺	1.0 ^a
CCl ⁺ + NH ₃ → HCl + HCNH ⁺	1.30 ^d
CCl ⁺ + H ₂ CO → H ₂ CCl ⁺ + CO	0.53 ^d
H ₂ CCl ⁺ + e ⁻ → Cl + CH ₂	300. ^a
H ₂ CCl ⁺ + C → HCl + C ₂ H ⁺	1.0 ^a
H ₂ CCl ⁺ + O ^b → HCl + HCO ⁺	1.0 ^a
H ₂ CCl ⁺ + H ₂ O → HCl + H ₃ CO ⁺	1.90 ^a

^a Guessed reaction and estimated rate constant.

^b Most significant reactions.

^c Bush *et al.* 1973.

^d Measured in this work.

^e Smith and Adams 1985.

^f Anicich *et al.* 1976.

^g Smith and Adams 1981.

TABLE 4
CONSTANTS AND DENSITY-DEPENDENT VARIABLES
IN THE CHLORINE MODEL

Constants	Density-dependent Variables
$\gamma = 1 \times 10^{-17} \text{ s}^{-1} \dots\dots$	$f_{\text{H}^+} \approx 5 \times 10^{-10}/n(f_{\text{O}} + f_{\text{H}_2\text{O}})$
$f_{\text{Cl}} = 1 \times 10^{-7} \dots\dots\dots$	$f_{\text{He}^+} \approx 2.5 \times 10^{-10}/n f_{\text{CO}}$
$f_{\text{O}} = 5 \times 10^{-5} \dots\dots\dots$	$f_{\text{O}^+} \approx 2 \times 10^{-11}/n f_{\text{H}_2\text{O}}$
$f_{\text{C}} = 5 \times 10^{-6} \dots\dots\dots$	$f_{\text{CH}_3^+} \approx 5 \times 10^{-11}/n f_{\text{O}}$
$f_{\text{N}} = 2 \times 10^{-6} \dots\dots\dots$	$f_{\text{H}_3^+} + f_{\text{HCO}^+} \approx 3 \times 10^{-12}/n f_e$
$f_{\text{CO}} = 5 \times 10^{-5} \dots\dots\dots$...
$f_{\text{H}_2\text{O}} = 2 \times 10^{-6} \dots\dots\dots$...
$f_{\text{NH}_3} = 1 \times 10^{-7} \dots\dots\dots$...
$f_{\text{H}_2\text{CO}} = 1 \times 10^{-8} \dots\dots\dots$...
$f_e = 1 \times 10^{-8} \dots\dots\dots$...

lar clouds (Prasad and Huntress 1980; Graedel, Langer, and Frerking 1982; Leung, Herbst, and Huebner 1984), while those for the density-dependent variables (ions) listed in Table 4 are calculated from equations incorporating the major production and destruction channels found from such theoretical models of oxygen-rich dense clouds ($10^3 < n < 10^6 \text{ cm}^{-3}$).

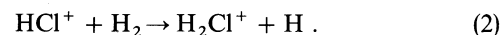
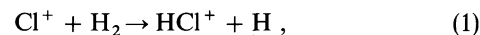
The model results for a cloud of "typical" composition with a density of $10^5 \text{ molecules cm}^{-3}$ are listed in Table 5 and are quite similar to that found for phosphorus (Thorne *et al.* 1984). A significant amount of atomic material remains, and two neutral species, Cl and HCl, dominate the interstellar chlorine chemistry, comprising ~99.9% of the gas-phase chlorine abundance. H_2Cl^+ and CCl^+ follow the neutral species in abundance, but are lower by at least a factor of 1000. The dominant production and loss channels for each species are indicated by note (b) in Table 3 and are illustrated in Figure 1, which summarizes the important chlorine chemistry in dense clouds, as judged by this reduced set of reactions. Again, none of the interstellar nonchlorine species are significantly affected by this chemistry because of the low fractional abundance of chlorine.

The prediction that HCl is a major component of the gas-phase chlorine abundance in dense clouds is in agreement with astronomical observations and shows that the species comprising interstellar chlorine differ significantly in diffuse and dense clouds. Several modifications to early chlorine models, which

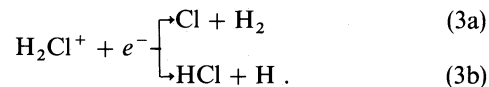
TABLE 5
CHLORINE MODEL RESULTS

Species	Fractional Abundance (for $f_{\Sigma\text{Cl}} = 10^{-7}$)
Cl.....	3.8×10^{-8}
HCl.....	6.2×10^{-8}
Cl^+	1.2×10^{-17}
HCl^+	3.9×10^{-17}
H_2Cl^+	3.3×10^{-12}
CCl^+	2.4×10^{-13}
H_2CCl^+	1.3×10^{-15}

produced too much HCl and too little Cl^+ , have been suggested to explain the lack of HCl in diffuse clouds. In these objects, neutral atomic chlorine is ionized by the ambient UV field and the resulting Cl^+ then produces the ionic precursor to hydrogen chloride, H_2Cl^+ , via H-atom abstraction reactions with H_2 :



The high electron abundance in diffuse clouds rapidly neutralizes H_2Cl^+ , which does not react with H_2 , producing Cl and HCl:



The major loss channels for HCl are photodissociation and reaction with C^+ . By including the C^+ channel, neglected in earlier models, increasing the HCl photodissociation rates, as has been suggested by recent *ab initio* calculations (van Dishoeck, van Hemert, and Dalgarno 1982), and using a small ($\leq 10\%$) branching ratio into channel (3b), van Dishoeck (1984) has been able to drive the predicted HCl abundance to below the derived observational upper limits while simultaneously increasing the Cl^+ abundance to near that observed if a slow temperature-dependent rate constant is used for reaction (1). Experimental measurements by Smith and Adams

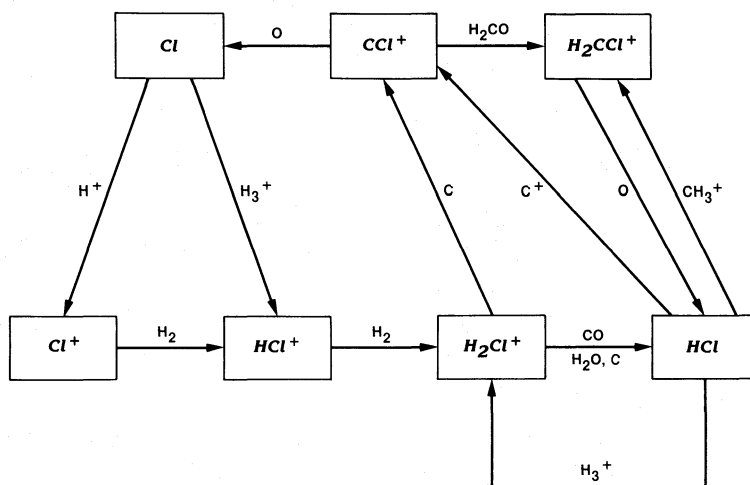
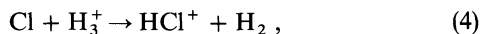


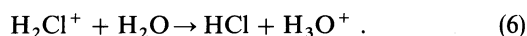
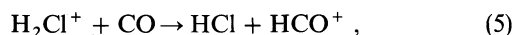
FIG. 1.—Summary of reactions that were found to be important in the dense cloud model. Arrows indicate reaction pathways: the species written next to the arrows indicate the other reactant in the bimolecular reaction.

(1981) to temperatures as low as 80 K indicate that both reactions (1) and (2) are reasonably temperature insensitive, but that reaction (1) is only mildly exothermic ($\Delta H \approx 5$ kcal per mole). A modest activation barrier cannot, therefore, be excluded at very low temperatures.

In dense clouds any atomic chlorine will remain predominantly neutral, so the formation of hydrogen chloride is initiated mainly by the reaction of Cl and H_3^+ :



followed by reactions (2) and (3). A slow reaction between Cl^+ and H_2 would therefore not limit the production of HCl in dense clouds. The proton affinity of HCl, which is slightly below that of CO, implies that H_2Cl^+ will react rapidly with abundant interstellar neutral species, such as CO and H_2O , to produce HCl:



Even though the difference between the proton affinities of HCl and CO is very small ($\lesssim 3$ kcal per mole), reaction (5) proceeds rapidly at 300 K (Table 1). Smith and Adams (1985) have also shown reaction (5) to proceed rapidly at lower temperatures, but that the reverse reaction is very slow. For the low predicted electron abundance [$f(e^-) \approx 10^{-8}$] and high molecular content [$f(\text{CO}) \approx 10^{-4}$] of dense interstellar clouds, reactions (5) and (6) proceed at significantly faster rates than the electron recombination processes listed in reactions (3a) and (3b), and the impact of any reduction of the branching ratio into channel (3b) is small. Thus, the removal of H_2Cl^+ by electrons in dense, dark clouds is orders of magnitude less efficient than in diffuse clouds, and the production rates of chlorine-bearing molecules are correspondingly higher. One might therefore naively expect that chlorine would be fully associated into HCl in dense clouds, since the major impediments to its production in diffuse clouds are no longer of importance. For the most part this is true, but several processes also act to limit the HCl fraction to $\sim 50\%$ of the available chlorine abundance.

The largest HCl loss channels involve reactions of H_2Cl^+ with C, and HCl with C^+ , both of which produce CCl^+ . CCl^+ is isoelectronic with CS and is quite stable. The IP of CCl^+ is low, and the CCl^+ ion is therefore relatively unreactive (see Table 1). Reaction with O, charge transfer to neutrals, and electron recombination are the most likely loss processes for CCl^+ in dense clouds. The net effect of these reactions, as shown in Figure 1, is to regenerate Cl and to establish a reaction pathway which cycles the chlorine rapidly between Cl and HCl. All other species in this pathway have abundances which are quite small with respect to the neutral endpoints. The steady state abundance ratio of Cl to HCl is fairly insensitive to order of magnitude variations in the chemical composition of the model molecular clouds, being most sensitive to the abundance of C^+ and the rate constant for the reaction of H_2Cl^+ with CO. The backward channel in reaction (5) is not important because proton transfer between HCl and H_3^+ occurs at a much faster rate, but the reaction of H_2Cl^+ and CO contributes substantially to the predicted HCl abundance. Even so, alteration of the C^+ abundance by factors of 10 and complete elimination of the proton transfer reaction between H_2Cl^+ and CO only change the estimated fractional abundance of HCl by at most a factor of 2, with the higher and lower C^+ abundances producing less or more HCl. Order of magnitude variations in

other important reactants such as C, O, and e^- produce even smaller changes in the predicted HCl abundance, on the order of 15% or less.

Astronomical observations and the predictions of large-scale modeling efforts suggest that the C and C^+ atoms will have significant abundances only for moderate density regions and that CO will dominate the carbon abundance in very dense cloud cores (Keene *et al.* 1985; Prasad and Huntress 1980; Leung, Herbst, and Huebner 1984; Tarafdar *et al.* 1985), implying that the abundance of HCl should increase with density. By incorporating these predictions into our model we find that the HCl abundance does increase with density, but that for densities above 10^5 this increase is not rapid. We therefore expect the fractional abundance of HCl in the central regions of giant molecular clouds to rise smoothly with density, with a fairly sharp increase in abundance where the column density in the outer cloud halo becomes sufficiently high such that photodissociation processes may be ignored, as is shown in Figure 2. For regions outside this photodissociation transition zone the remaining chlorine will be a mixture of Cl and Cl^+ , while inside this zone the atomic chlorine will remain predominantly neutral. Under all of these conditions, the predicted fraction of gas-phase chlorine present as HCl varies between 25%–65% in dense clouds. This produces an estimated fractional abundance of $X(\text{HCl}) \approx 2.5\text{--}6.5 \times 10^{-8}$ for a depletion factor of 3, in good agreement with the observational estimates of $0.5\text{--}5.0 \times 10^{-8}$ (Blake, Keene, and Phillips 1985).

IV. CONCLUSIONS

Simple kinetic modeling of interstellar chlorine chemistry shows that reactions of $H_3^+ + \text{Cl}$ and $H_2 + HCl^+$, in combination with rapid proton transfer reactions involving H_2Cl^+ , produce substantial amounts of HCl in dense interstellar clouds in agreement with astronomical observations. The kinetic model also predicts that reactions of $H_2Cl^+ + \text{C}$ and $\text{HCl} + C^+$ will regenerate atomic chlorine, leading to a chemical cycle in which Cl and HCl both retain significant gas-phase abundances. The predictions that Cl is also a major repository of chlorine in dense clouds and that Cl and HCl by themselves constitute the overwhelming fraction of chlorine in such objects may be tested by observing the fine-structure transition of atomic chlorine at $11.3 \mu\text{m}$ in absorption toward such objects as Orion. The frequency of this transition is known to better than 0.3 km s^{-1} from infrared laser magnetic resonance (LMR) work (Dagenais, Johns, and McKellar 1976), and calculations indicate that for the suggested Cl abundance, the optical depth of the fine-structure transition should approach or exceed unity, making its detection feasible with high-resolution instruments.

Finally, the chlorine depletion in dense interstellar clouds may be estimated by combining the calculated model HCl fractional abundances with limits inferred from astronomical observations of the $J = 1 \rightarrow 0$ transition. For a cosmic chlorine abundance of 3×10^{-7} , our results indicate that the chlorine depletion in Orion lies between 3–30. The large uncertainty arises mainly from difficulties associated with the astronomical observations and their interpretation. As mentioned earlier, Jura and York (1978) find from optical studies that chlorine in diffuse clouds is depleted by no more than a factor of 3, although more recent work by Tarafdar, Prasad, and Huntress (1983) and Harris, Gry, and Bromage (1984) argues that elemental depletion in clouds increases strongly with either H_2

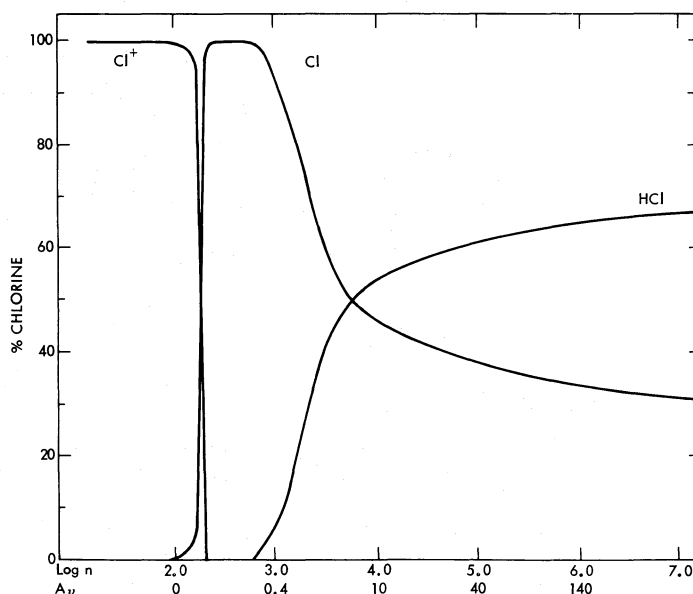


FIG. 2.—Variation of the Cl/HCl abundance with density (A_v). The A_v scale has been established using the collapsing cloud model of Tarafdar *et al.* (1985) for a cloud of total mass $1000 M_\odot$.

column density or number density and that chlorine may be depleted by a factor of 5 or more. In any case, the estimated chlorine depletion in dense clouds appears to be at most a factor of 10 above that in diffuse clouds, which implies that, compared with other more refractory second-row elements like Si and P, gas-phase chlorine species may well remain abundant

even in the cores of dense molecular clouds.

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