

Nitrogen-15 Magnetic Resonance Spectroscopy. VII. Chemical Shifts of *para*-Substituted Nitrobenzenes¹

David T. Clark and John D. Roberts

Contribution No. 3308 from the Gates and Crellin Laboratories of Chemistry,
California Institute of Technology, Pasadena, California.

Received September 30, 1965

Abstract: The ¹⁵N chemical shifts of a series of *para*-substituted nitrobenzenes have been measured. Comparison with the *para*-¹³C chemical shifts in monosubstituted benzenes and with ¹⁹F chemical shifts in *para*-substituted fluorobenzenes, indicates that similar factors determine the chemical shifts of these three nuclei. With *p*-nitroaniline, the ¹⁵N chemical shift appears to have an extra paramagnetic contribution, since this is the only compound studied that does not fit the correlation with ¹³C and ¹⁹F chemical shifts.

In the Pople² theory of chemical shifts, the screening constant of a nucleus A is divided into four contributions

$$\sigma_A = \sigma_{AA}^d + \sigma_{AA}^p + \sum_{A \neq B} \sigma_{AB} + \sigma_{A, \text{ring}}$$

where σ_{AA}^d and σ_{AA}^p are the diamagnetic and paramagnetic contributions from the electrons of atom A, σ_{AB} is the contribution from the electrons of atom B, and $\sigma_{A, \text{ring}}$ is the contribution due to ring currents. Investigations in this and other laboratories have shown that in the case of carbon,^{3,4} nitrogen,⁵⁻⁸ and fluorine,^{9,10} the main contributions to the shielding constant of the nucleus A comes from the paramagnetic term σ_{AA}^p . If the comparison of shielding parameters is confined to a suitable series of closely related compounds, e.g., substituted benzene derivatives, then contributions from the substituent anisotropy ($\sum_{A \neq B} \sigma_{AB}$) and ring current ($\sigma_{A, \text{ring}}$) terms may be expected to remain essentially constant. If the mean electronic excitation energies, ΔE , involved in the closure approximation for the σ_{AA}^p 's are also constant, then the contribution to the shielding constants from σ_{AA}^d and σ_{AA}^p may be related in a simple way to the local electron density at the nucleus. Such correlations have been reported for ¹³C chemical shifts^{11,12} in monosubstituted benzenes and ¹⁹F chemical shifts^{10,12} in *para*-substituted fluorobenzenes.

By contrast there has been no detailed investigation of ¹⁵N chemical shifts in an analogous series of compounds despite the obvious theoretical interest of such an investigation. With this in mind we have examined ¹⁵N chemical shifts of a series of *para*-substituted ¹⁵N nitrobenzenes in the hope of demonstrating similar correlations to those found previously for the ¹³C and ¹⁹F chemical shifts.

(1) Supported in part by the Public Health Service Research Grant 11072-02 from the Division of General Medical Sciences and the National Science Foundation.

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(11) H. Spiesscke and W. G. Schneider, *ibid.*, **35**, 731 (1961).

(12) T. K. Wu and B. P. Dailey, *ibid.*, **41**, 2796 (1964).

The series of nitro compounds were chosen for study for two reasons: firstly, convenience of preparation, and secondly, because the main contribution to the paramagnetic term (σ_{AA}^p) is expected to arise from the $\sigma \rightarrow \pi^*$ transitions of the nitrogen atoms.¹³ The energy of the latter transitions should remain roughly constant throughout the series. Also the variation in the ring current ($\sigma_{A, \text{ring}}$) and anisotropy terms ($\sum_{A \neq B} \sigma_{AB}$) are expected to remain constant throughout the series. (The main contribution to $\sum_{A \neq B} \sigma_{AB}$ arises from the $n \rightarrow \pi^*$ transition of the oxygen lone pairs, and the energy of this transition is little affected by substituents in the *para* position, e.g., nitrobenzene $n_0 \rightarrow \pi^*$ 330 m μ ,¹⁴ *p*-nitrophenol $n_0 \rightarrow \pi^*$ 334 m μ .¹⁵)

Results and Discussion

Table I gives the ¹⁵N chemical-shift differences of a series of *para*-substituted nitrobenzenes relative to ¹⁵N nitrobenzene as 15 mole % solutions in dimethyl sulfoxide. Bulk susceptibility corrections were not applied, since these are likely to be well within the experimental error (± 0.16 p.p.m.), for the relatively dilute solutions employed. The low sensitivity (10^{-3} that of ¹H at constant field) precludes any accurate measurement of chemical shifts over a wide range of concentrations.

The first point to note from Table I is the relatively small range of chemical shifts, namely 8.23 p.p.m. This may be compared with the corresponding ¹⁹F and ¹³C chemical shifts in *para*-substituted fluorobenzenes of 24.35 p.p.m.^{16a} and ¹³C chemical shifts in *para*-substituted toluenes of 1 p.p.m.^{16b} for the same substituents. The intermediate range of the ¹⁵N chemical shifts is as expected if the paramagnetic terms σ_{AA}^p dominate the shielding constants for these nuclei, and is indicative that the mean electronic excitation energies are at least approximately constant in the series.¹⁷

Previous investigations have correlated the ¹³C chemical shifts of the *para*-carbon atom in monosubstituted benzenes with the ¹⁹F chemical shifts of *para*-

(13) See ref. 3, p. 2805.

(14) K. L. Wolf and W. Herold, *Z. Physik. Chem.*, **B13**, 201 (1931).

(15) K. Semba, *Bull. Chem. Soc. Japan*, **33**, 1640 (1960).

(16) (a) R. W. Taft, E. Price, I. R. Fox, I. C. Lewis, K. K. Anderson, and G. T. Davis, *J. Am. Chem. Soc.*, **85**, 3146 (1963); (b) P. C. Lauterber, *Ann. N. Y. Acad. Sci.*, **70**, 841 (1958).

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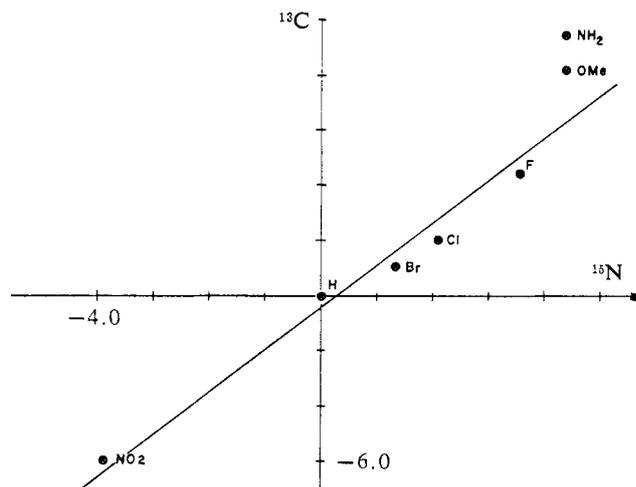


Figure 1. Plot of *para*- ^{13}C chemical shifts in monosubstituted benzenes vs. ^{15}N chemical shift of nitro group in *para*-substituted nitrobenzenes. Data for ^{13}C are from ref. 11.

substituted fluorobenzenes.¹² In the former case, the chemical shifts are determined by the π -electron distributions at the carbon atoms,¹¹ and the implied correlation of ^{19}F chemical shifts with π -electron density at the attached carbon atom then arises out of the near equivalence of the terms involving the C-F π bond order and the π -electron density on fluorine.¹⁰

Table I. ^{15}N Chemical Shifts of the Nitro Group in *para*-Substituted Nitrobenzenes

Substituent	Chemical shift, p.p.m. ^a
NH ₂	4.38
OMe	4.38
F	3.57
NHCOCH ₃	3.07
Cl	2.09
Br	1.32
H	0.00
CN	-2.58
NO ₂	-3.85

^a Experimental error is approximately ± 0.16 p.p.m. A positive value of the chemical shift indicates a shift upfield from nitrobenzene.

The ^{15}N chemical shifts are shown plotted against the *para*- ^{13}C chemical shifts in monosubstituted benzenes and ^{19}F chemical shifts in *para*-substituted fluorobenzenes in Figures 1 and 2. There is a good over-all correlation in the two graphs except possibly for the ^{15}N chemical shift in *p*-nitroaniline. This would seem to confirm that similar factors are operative in determining the chemical shift of ^{15}N , ^{19}F , and ^{13}C nuclei.

By analogy with the theory developed for the corresponding fluorine chemical shifts,¹⁸ the ^{15}N nitro chemical shift differences in *para*-substituted nitrobenzenes, should depend on the following factors: (1) the π -electron density of the nitrogen atom; (2) the π -electron density at the adjacent carbon and oxygen atoms; (3) the N-C π bond order; (4) the N-O π bond order. The presence of the two oxygen atoms in

(18) R. W. Taft, F. Prosser, L. Goodman, and G. T. Davis, *J. Chem. Phys.*, **38**, 380 (1963).

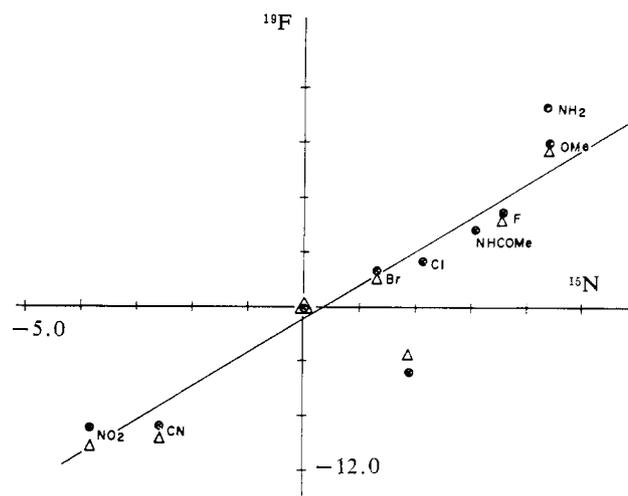


Figure 2. Plot of ^{19}F chemical shifts in *para*-substituted fluorobenzenes vs. ^{15}N chemical shift of nitro group in *para*-substituted nitrobenzenes: Δ , ^{19}F chemical shifts in dimethyl sulfoxide; \bullet , ^{19}F chemical shifts in cyclohexane.

the nitro group means that the expression for the shielding parameter of the nitrogen atom will be considerably more complex than that for fluorine. However, whatever the contributing factors are, they are not all expected to operate in the same direction.¹⁰ For example, a relatively strong mesomeric substituent, such as methoxyl, should increase the electron density at the nitrogen atom (1), and also at the carbon and oxygen atoms (2), thus resulting in an upfield shift relative to nitrobenzene. The same substituent is expected to increase the N-C π bond order (3), while the N-O π bond order should decrease (4), the first term causing a downfield shift, the latter an upfield shift. Despite the possible subtle interplay of these factors, there is a good correlation with the corresponding ^{19}F and ^{13}C chemical shifts, only *p*-nitroaniline is possibly anomalous. The amino group is known from spectroscopic data to have a much larger (+M) effect than the other substituents,¹⁹ and this and dipole moment data²⁰ indicate that there is considerable interaction between the amino and nitro groups. Since the increase in C-N π bond order is expected to be larger than the decrease in N-O π bond order, the extra paramagnetic shift implies that the former term more than compensates for the increased electron densities at the carbon, nitrogen, and oxygen atoms in *p*-nitroaniline. An alternative explanation would be that for *p*-nitroaniline, the charge-transfer states are much lower in energy than for the other nitrobenzenes.²¹ The mean electronic excitation energy ΔE for the $\sigma \rightarrow \pi^*$ transition might thus be expected to be smaller in this case, leading to an increased paramagnetic (σ^{PAA}) contribution to the shielding constant

Experimental Section

The nitrogen-15 spectra were obtained with a Varian Model V-4300B spectrometer, operated at 6.08 Mc.p.s. and 14,100 gauss,

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(20) A. L. McClellan, "Tables of Experimental Dipole Moments," W. H. Freeman and Co., San Francisco, Calif., 1963, p. 192.

(21) Cf. ref. 20, Chapter 10.

as described previously.^{7,22} Most of the following compounds were prepared on a 0.5-g. scale.

p-Nitro-¹⁵N-acetanilide was prepared by the nitration of acetanilide with 38% nitric ¹⁵N acid (Merck Sharp and Dohme 97% ¹⁵N), according to the procedures given by Vogel.²³

p-Nitro-¹⁵N-aniline was obtained by the acid hydrolysis of *p*-nitro-¹⁵N-acetanilide.²³

Nitrobenzene-¹⁵N was prepared by the nitration of benzene with 38% nitric-¹⁵N acid according to the method given by Vogel.²⁴

p-Bromonitrobenzene-¹⁵N was prepared by the nitration of bromobenzene with 38% nitric-¹⁵N acid according to the method given by Vogel.²⁵

p-Chloronitrobenzene-¹⁵N was prepared by the nitration of chlorobenzene with 38% nitric-¹⁵N acid using similar conditions to those in the *p*-bromonitrobenzene-¹⁵N preparation.

(22) G. Binsch, J. B. Lambert, B. W. Roberts, and J. D. Roberts, *J. Am. Chem. Soc.*, **86**, 5564 (1964).

(23) A. I. Vogel, "A Textbook of Practical Organic Chemistry," 3rd Ed., Longmans, Green and Co., London, p. 581.

(24) *Cf.* ref. 23, p. 525.

p-Fluoronitrobenzene-¹⁵N was prepared from fluorobenzene and 38% nitric-¹⁵N acid in the presence of sulfuric acid at 70°. The product was isolated in the same way as for the nitrobenzene-¹⁵N preparation.

p-Nitro-¹⁵N-anisole was prepared by the reaction of *p*-bromonitro-¹⁵N-benzene with sodium methoxide in dimethyl sulfoxide solution at room temperature. The reaction mixture was poured onto crushed ice, and the pale yellow crystals of nitro-¹⁵N-anisole were removed by filtration. The yield was almost quantitative.

p-Dinitro-¹⁵N-benzene (labeled with ¹⁵N in one nitrogen) was prepared by diazotized *p*-nitro-¹⁵N-aniline according to the procedure given by Hodgson, *et al.*²⁶

p-Nitro-¹⁵N-benzonitrile was prepared from *p*-nitro-¹⁵N-aniline by the Sandmeyer reaction in a manner analogous to that previously described by Bogert and Hand.²⁷

(25) *Cf.* ref. 23, p. 527.

(26) H. H. Hodgson, F. Heyworth, and E. R. Ward, *J. Chem. Soc.*, 1512 (1948).

(27) M. T. Bogert and W. F. Hand, *J. Am. Chem. Soc.*, **24**, 1035 (1902).

Neighboring Carboxyl Group Participation in the Hydrolysis of Monoesters of Phthalic Acid. The Dependence of Mechanisms on Leaving Group Tendencies

John W. Thanassi¹ and Thomas C. Bruice²

Contribution from the Departments of Biological Sciences and Chemistry, University of California at Santa Barbara, Santa Barbara, California.

Received October 9, 1965

Abstract: The hydrolyses of the following monohydrogen phthalate esters have been investigated in the region of the carboxyl group pK_a' : I, methyl; II, 2'-monochloroethyl; III, propargyl; IV, 2',2',2'-trifluoroethyl; and V, phenyl. In addition, the hydrolysis of (VI) O-phthaloyl-N-acetylserinamide was also investigated. The pK_a' values of the alcohols employed in the preparation of the esters fall in the order I > II > III = VI > IV > V. Esters I and II hydrolyzed with neighboring COOH group participation while esters IV and V hydrolyze with COO⁻ participation. Thus, the mechanism changes with increase in leaving tendencies. The crossover between mechanisms occurs for esters of alcohols with pK_a' values of *ca.* 13.5 so that esters III and IV are equally prone to COOH and COO⁻ neighboring group catalyzed hydrolysis. These results are considered in the light of the possible participation of carboxyl groups in the mechanism of esterases. Extrapolation from this study suggests that such participation could be effective but defies detection through conventional examination of pH-rate profiles.

In order to understand the mechanism of enzyme action, it is important that the functional groups on the protein necessary for activity be identified and that their mode of action in the catalytic process be elucidated. The study of intramolecular catalysis in model systems has been a most useful tool in this direction.³ This paper deals with neighboring carboxyl group participation in the hydrolysis of half-esters of phthalic acid.

It has been demonstrated that carboxyl anion participates in the hydrolyses of acetyl salicylic acid,⁴ substituted phenyl acid succinates and glutarates,⁵ and mono-*p*-bromophenyl esters of substituted glutaric and succinic acids.⁶ On the other hand, ethyl hydrogen

phthalate⁷ and methyl hydrogen 3,6-dimethylphthalate⁸ hydrolyze with participation by the neighboring protonated carboxyl group. It would appear, then, that those monoesters of dicarboxylic acids having good leaving groups hydrolyze with assistance by neighboring carboxyl anion (COO⁻) and those having poor leaving groups hydrolyze with the aid of the neighboring carboxyl group (COOH).

Carboxyl groups have been implicated in the mechanism of action of several enzymes. Thus, Bernhard and Gutfreund⁹ have found that in the ficin-catalyzed hydrolysis of benzoyl arginine ethyl ester, a carboxyl anion apparently participates in the hydrolysis of the acyl enzyme. This has been interpreted to occur by means of a rate-limiting attack of the carboxyl anion

(1) Postdoctoral Fellow, University of California at Santa Barbara.

(2) To whom inquiries should be addressed.

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