

## Nuclear Magnetic Resonance Spectroscopy. A Stereospecific $^3J_{CF}$ Coupling in the Low-Temperature $^{13}\text{C}$ Nmr Spectrum of 1,1-Difluorocyclohexane\*

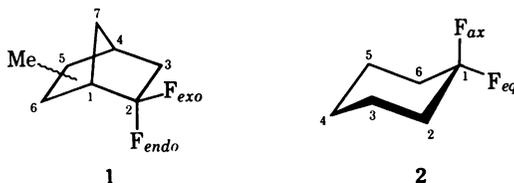
David Doddrell, Claude Charrier, and John D. Roberts

CONTRIBUTION NO. 4115 FROM THE GATES AND CRELLIN LABORATORIES OF CHEMISTRY,  
CALIFORNIA INSTITUTE OF TECHNOLOGY, PASADENA, CALIF. 91109

Communicated August 10, 1970

**Abstract.** The proton-decoupled  $^{13}\text{C}$  nmr spectrum of 1,1-difluorocyclohexane has been examined at room temperature and at  $-90^\circ\text{C}$ . There are only minor changes in the one-bond and two-bond carbon-fluorine scalar coupling constants at the lower temperature; however, the triplet observed for C-3 ( $^3J_{CF} = 4.7$  Hz) collapses to a doublet ( $^3J_{CF} = 9.5$  Hz) at  $-90^\circ\text{C}$ . It is proposed that only the equatorial fluorine is coupled with the C-3 carbon as the result of operation of a back-lobe orbital interaction.

In a previous publication<sup>1</sup> concerning the  $^{13}\text{C}$  proton-decoupled nmr spectra of norbornyl derivatives, a number of methyl-substituted 2,2-difluoronorbornanes (**1**) were examined. It was found that the C-7 resonance appeared as a doublet with  $^3J_{CF}$  on the order of 5 Hz. It was proposed that, in these compounds, carbon-fluorine coupling occurs only with the *endo*-fluorine. This contention was supported by the  $^{13}\text{C}$  spectrum of *exo*-2-fluoro-norbornane in which C-7 appeared as a sharp singlet.



Similar behavior might be expected for 1,1-difluorocyclohexane (**2**), provided that the chair-chair interconversion could be halted. The molecular dynamics of **2** are well known,<sup>2</sup> and at  $-90^\circ\text{C}$  the chair-chair interconversion is slow on the nmr time scale. At room temperature (Fig. 1) the  $^{13}\text{C}$  spectrum, with chemical shifts in ppm upfield from internal  $\text{CS}_2$  shows triplets at 69.9,<sup>3</sup> ( $^1J_{CF} = 242$  Hz); 165.3, ( $^2J_{CF} = 24.0$  Hz); and 175.5, ( $^3J_{CF} = 4.7$  Hz) assigned to C-1, C-2, and C-3, respectively. The resonance of C-4 appeared as a broad singlet at 173.6 ppm. At  $-90^\circ\text{C}$ , (Fig. 2), there were only minor ( $<1$  Hz) changes in the  $^1J_{CF}$  and  $^2J_{CF}$  values; however, the resonance of C-3 now appeared as a doublet,  $^3J_{CF} = 9.5$  Hz. Note that the room-temperature value of  $^3J_{CF}$  is half the value observed in the low-temperature limit; we believe, by analogy with the results<sup>1</sup>

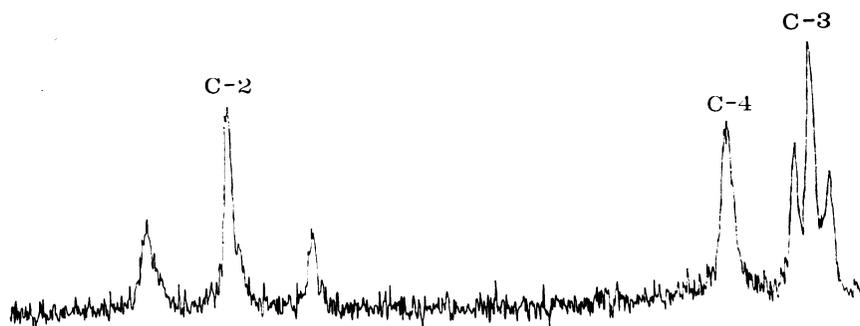


FIG. 1. Proton noise-decoupled  $^{13}\text{C}$  spectrum of 1,1-difluorocyclohexane at room temperature. The resonance for C-1 is not shown.

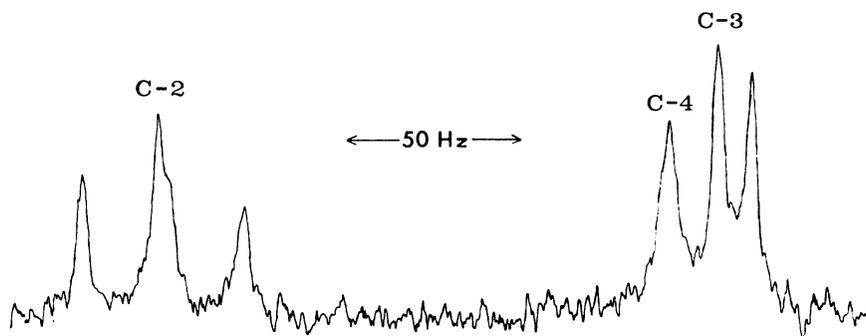


FIG. 2. Proton noise-decoupled  $^{13}\text{C}$  spectrum of 1,1-difluorocyclohexane at  $-90^\circ\text{C}$ . The resonance for C-1 is not shown.

observed for **1**, that C-3 is coupled only to the equatorial fluorine. This is consistent with the recent theory,<sup>3,4</sup> involving the coupling between heavy nuclei, that has proposed that there is a significant contribution from through-space interactions (presumably via back-lobe orbital overlap).

It has been proposed<sup>5</sup> that the weak  $^{13}\text{C}$  spin-echo Fourier transform (SEFT) spectrum of cyclohexane is due to a high carbon-13  $T_1/T_2$  ratio, resulting from an extra contribution to the carbon-13  $T_2$  relaxation time by modulation of the scalar C, H coupling constant associated with the chair-chair interconversion. It was not stated which C, H coupling was being modulated. We believe that the results obtained for **2** are evidence of a significant modulation of  $^3J_{\text{CH}}$  in cyclohexane and to a short carbon-13  $T_2$  value. We would predict a similarly weak spin-echo Fourier transform spectrum for perfluorocyclohexane.

\* Supported by the National Science Foundation.

<sup>1</sup> Grutzner, J. B., M. Jautelat, J. B. Dence, R. A. Smith, and J. D. Roberts, *J. Amer. Chem. Soc.*, in press.

<sup>2</sup> Spassov, S. L., D. L. Griffith, E. S. Glazer, K. Nagarjan, and J. D. Roberts, *J. Amer. Chem. Soc.*, **89**, 88 (1967).

<sup>3</sup> Barfield, M., *J. Chem. Phys.*, **41**, 3825 (1964).

<sup>4</sup> Karplus, M., and M. Barfield, *J. Amer. Chem. Soc.*, **91**, 1 (1969).

<sup>5</sup> Allerhand, A. and D. W. Cochran, *J. Amer. Chem. Soc.*, **92**, 4482 (1970).