

Zintl phases for thermoelectric devices

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By converting waste heat into electricity and improving the efficiency of refrigeration systems, thermoelectric devices could play a significant role in solving today's energy problems. Increasing the thermoelectric efficiency (as measured by the thermoelectric material's figure-of-merit, zT) is critical to the development of this technology. Complex Zintl phases, in particular, make ideal candidates for thermoelectric materials because the necessary "electron-crystal, phonon-glass" properties can be engineered with an understanding of the Zintl chemistry. A recent example is the discovery that $\text{Yb}_{14}\text{MnSb}_{11}$, a transition metal Zintl compound, has twice the zT as the material currently in use at NASA. This perspective outlines a strategy to discover new high zT materials in Zintl phases, and presents results pointing towards the success of this approach.

1 Introduction

In recent years there has been a renewed effort to identify high performance thermoelectric materials for applications in power generation and device cooling. Several factors have driven this interest, including environmental concern with the use of chlorofluorocarbons in compressor-based refrigeration and the potential to generate electrical power from low-level waste heat. In particular, thermoelectric power generation is being considered for use in automobiles; here, the waste exhaust heat can be utilized, thereby improving the overall fuel economy of the vehicle. Fig. 1 shows a schematic of a thermoelectric device. Heat from above drives the free carriers (electrons in the n-type leg, holes in the p-type leg) from the hot side of the thermoelectric material to the cold side, thereby creating electrical power in the load. Alternatively, thermoelectrics can be used for refrigeration because replacing the load with a voltage source will force heat to flow from the top side to the bottom side. Because they have no moving parts and use no liquid refrigerant, thermoelectric refrigerators and power generators are particularly well-suited to applications in which long-term, unattended operation is necessary, or very small devices are to be cooled.¹ Current applications include power generation for deep-space probes or remote sensors, and thermal management of high power electronics. While many applications of thermoelectric devices can be envisioned, their widespread implementation is limited as a result of the low efficiency that today's thermoelectric materials afford.

The efficiency of a thermoelectric device generally improves by increasing the material's thermoelectric figure of merit $zT = a^2 T / \rho\kappa$ (a : Seebeck coefficient (V K^{-1}), T : temperature (K), ρ : electrical resistivity ($\Omega \text{ cm}$), κ : thermal conductivity ($\text{W cm}^{-1} \text{ K}^{-1}$)).^{1,2} For metals or degenerate semiconductors the Seebeck coefficient is proportional to temperature, effective mass m^* , and inversely proportional to the charge carrier concentration n . In the approximation of energy independent scattering, the relationship is

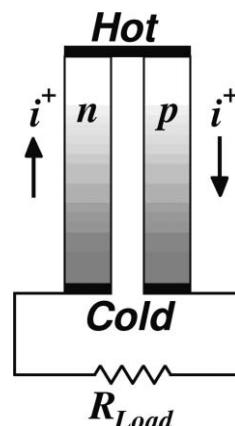


Fig. 1 Schematic of a thermoelectric generator.

given³ by

$$a = \frac{8\pi^2 k_B^2}{3e h^2} m^* T \left(\frac{\pi}{3n} \right)^{2/3} \quad (1)$$

where k_B is Boltzmann's constant, e is the charge of an electron and h is Planck's constant. Thus, large thermopower $|a|$ translates to large effective mass which is equivalent to large density-of-states or flat bands at the Fermi level. Low electrical resistivity (or high electrical conductivity σ) requires a large mobility, μ , and large carrier (electrons or holes) concentration n through

$$1/\rho = \sigma = ne\mu \quad (2)$$

High zT requires a combination of the large thermopower found in insulators or semiconductors with the low electrical resistivity of metals. This is typically only achieved in heavily doped semiconductors combining high electron density-of-states (high effective masses) with high carrier concentrations, and high mobility¹ as seen in eqn (1) and (2) for a and σ .

The thermal conductivity κ can be separated into two parts: the lattice or phonon contribution κ_l and the electronic contribution κ_e , which is proportional to the electrical conductivity $\kappa_e = L\sigma T$ (with the Lorenz number $L = \pi^2 k_B^2 / 3e^2$). Thus in addition to optimizing the electronic transport properties, the lattice thermal

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Shawna R. Brown received her Bachelor's of Science in chemistry in 2001 from the University of California, Davis. She continued her education at UC Davis and achieved a Master's degree in inorganic chemistry in 2005 while working under the supervision of Professor Susan M. Kauzlarich. Her thesis titled, Thermoelectric Investigation of $Yb_{14}MSb_{11}$ ($M = Mn, Zn$) and $A_{11}Sb_{10}$ ($A = Ca, Yb$), focuses on examining the thermoelectric properties of Zintl compounds. Shawna's graduate research unearthed a novel thermoelectric material ($Yb_{14}MnSb_{11}$) for which she was recognized by the National Aeronautics and Space Administration. As a result of the exceptional properties of $Yb_{14}MnSb_{11}$ and related compounds, she has formed a start-up company, High Merit Thermoelectrics, Inc., that specializes in thermoelectric materials and devices. To date, her company has secured patenting and licensing rights for $Yb_{14}MnSb_{11}$, and is currently seeking funding to initiate their prototyping stage. In the future she aims to continue to develop improved thermoelectric materials for power generation applications.

G. Jeffrey Snyder obtained his B.S. degree in physics, chemistry and mathematics at Cornell University (1991) focusing on solid state chemistry of chalcogenides under the direction of Professor Frank DiSalvo. During 1992–1993 he studied the solid state chemistry of intermetallics, sub-nitrides and C_{60} materials with Professor Arndt Simon at the Max Planck Institut für Festkörperforschung in Stuttgart, Germany. He received his Ph.D. in applied physics from Stanford University (1997) where he was a Hertz Fellow. At Stanford he studied the magnetic and magneto-electrical transport properties of metallic perovskites under the direction of Professor Ted Geballe and Professor Mac Beasley. He was a Senior Member of the Technical Staff in the thermoelectrics group at NASA's Jet Propulsion Laboratory for 9 years (1997–2006) where he focused on thermoelectric materials and devices. Here he developed the thermoelectric compatibility concept for efficient segmentation of thermoelectric materials, electrochemical MEMS fabrication of thermoelectric microdevices, and a number of new, high efficiency thermoelectric materials. He is currently a Faculty Associate in materials science at the California Institute of Technology (Caltech). His recent interests include the discovery of new Zintl phase thermoelectric materials and nanostructured thermoelectric composites using bulk processing. Dr Snyder has published over 80 articles, and a number of patents including a recent chapter in the CRC Handbook of Thermoelectrics.



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conductivity must also be low, such as that of an amorphous glass, to achieve high zT . These criteria have led to the concept² that the ideal material has the electronic structure of a compound semiconductor with the atomic structure of a glass, or an “electron-

crystal, phonon–glass”. Finally because zT is very sensitive to carrier concentration (*via* a and ρ), it is essential to be able to control the carrier concentration through precise doping without disrupting the carrier mobility.

Today's commercial thermoelectric materials are alloys discovered decades ago. The alloying of isoelectronic species (such as Si–Ge alloys or Bi_2Te_3 – Sb_2Te_3 – Bi_2Se_3 alloys) introduces disorder which scatters phonons without too much disruption of the electron transport. While alloying substantially reduces lattice thermal conductivity, other mechanisms for reducing thermal conductivity are possible. Recent discoveries of high zT materials in oxides,⁴ nanostructured materials,^{5,6} skutterudites, clathrates, chalcogenides,⁷ and Zintl phases⁸ indicate that there are many fruitful systems yet to consider.

For room-temperature applications, such as refrigeration and waste heat recovery up to 475 K, Bi_2Te_3 and its alloys have proven to possess the greatest figure of merit for both n- and p-type thermoelectric systems. Bi_2Te_3 was first investigated as being a material of great thermoelectric promise in 1954.⁹ Today, more than five decades later, this compound semiconductor has remained the best commercial material for use near room temperature with maximum reported zT typically in the range of 0.8 to 1.1 for optimized systems.¹

For higher temperature power generation, materials based on Group 14 tellurides are typically employed such as PbTe , GeTe or SnTe . Reported zT 's are usually less than 0.8. However, alloys, particularly with AgSbTe_2 , have lead to several reports of $zT > 1$ for both n-type and p-type materials.^{10–13} The p-type alloy ($\text{GeTe})_{0.85}(\text{AgSbTe}_2)_{0.15}$ (TAGS) with maximum zT greater than 1.2 has been successfully used in long-life thermoelectric generators.

2 Zintl phases

The combination of requirements inherent in zT includes thermodynamic as well as both electrical and thermal transport properties (all of which can vary by orders of magnitude) which makes designing a thermoelectric material quite complex. One method to achieve this unusual mix of properties is to envision a complex material with distinct components providing different functions. This ideal thermoelectric material would have regions of the structure composed of a high mobility compound semiconductor that provides the “electron–crystal” electronic structure, interwoven (on a Å or nm scale) with a phonon inhibiting structure to act as the “phonon–glass”.² The phonon–glass region would benefit from disorder and therefore would be ideal to house dopants without disrupting the electron–crystal region.

Zintl phases are prime candidates for applying this concept to obtain high zT thermoelectric materials. Zintl phases are made up of electropositive cations (typically, Group 1 and 2) which donate their electrons to electronegative anions; which, in turn, use them to form bonds in order to satisfy valence.¹⁰ The structural requirements of Zintl phases are explained by assuming the presence of both ionic and covalent contributions to the bonding picture of the structure. Classical Zintl phases are considered to be valence precise semiconductors.¹¹ At the border between classical Zintl phases and normal metallic phases, the semiconducting band gap diminishes and metallic conductivity can result.¹² Metallic phases at this Zintl border are referred to as polar intermetallics.¹³ The simple Zintl formulism can apply to polar intermetallics, but the difference in electronegativity between the cationic and the anionic units is small, giving rise to homometallic and heterometallic bonding.^{16–18} Thus, the term

polar intermetallics is more appropriate when emphasizing that the extent of charge transfer between the cations and anions is significantly less than the ionic model, such as is the case in lithium containing Zintl phases.¹¹ Bonding within the polyanionic framework using the Zintl formulism is considered localized, 2-center–2-electron bonding, whereas in polar intermetallics charge transfer from the cation to the polyanionic unit is incomplete, and delocalized multicenter bonding is included.^{16–18} The term transition metal Zintl phase is another distinction used in the literature to describe transition metal containing compounds that are isostructural to main group metalloid containing Zintl compounds^{14–16} or, in some cases, new transition metal containing structures where the Zintl formulism provides insight into the electronic structure and bonding.^{22–26}

In a Zintl phase the Zintl anions (or transition metal–Zintl ions) provide the “electron–crystal” electronic structure through the covalently bonded network of complex anions or metalloids. Fig. 2 shows several examples of antimony containing Zintl phases: $\text{Ca}_{11}\text{GaSb}_9$,¹⁷ CaZn_2Sb_2 ,¹⁸ and BaGa_2Sb_2 .¹⁹ These compounds are all semiconductors and the bonding can be described via the Zintl formulism as consisting of electropositive cations and complex anions ($\text{Ca}_{11}\text{GaSb}_9$), a layered net (CaZn_2Sb_2) or an open framework or tunnel structure (BaGa_2Sb_2). $\text{Ca}_{11}\text{GaSb}_9$ consists of 11 Ca^{2+} cations which donate their electrons to GaSb_4^{9-} tetrahedron, Sb_2^{4-} dumbbell, and 3 Sb^{3-} isolated anions. The Sb anions are considered isolated since there are no near neighbor Sb atoms that can be considered bonding. These anions are 6 coordinated by the alkaline earth cation. One might expect low lattice thermal conductivity with this complex structure. The CaZn_2Sb_2 structure is described in detail below. BaGa_2Sb_2 is made up of Ba^{2+} cations within a $(\text{Ga}_2\text{Sb}_2)^{2-}$ network that is composed of $(\text{Ga}_2\text{Sb}_{6/3})^{2-}$ units which can be described as a Ga–Ga ethane-like staggered conformation with the 6 Sb atoms. These $\text{Ga}_2\text{Sb}_{6/3}$ units are linked in a way that leaves holes in the structure as shown in Fig. 2. Such an electron precise compound will naturally form bonding and antibonding bands that will create the necessary semiconducting band-gap. The covalent character of the polyanion bonding can provide high mobility to the charge carriers. Because the band gap arises due to the separation of bonding and antibonding states in the anions, the existence of nonbonding or partially bonding states can fill the gap region with states and make the compound metallic. In contrast, more electronegative elements (such as oxygen and sulfur) will have ionic character to their bonding which will trap carriers and reduce mobility. Zintl anions made from heavier, softer atoms will naturally have lower lattice thermal conductivity. In addition these heavy atoms make smaller band gap semiconductors that are more amenable to the high dopant concentrations necessary for thermoelectrics. The cations in Zintl phases provide regions that can be alloyed or precisely doped to control electron concentration and disrupt phonon transport through alloy scattering. In this way, the thermal and electronic properties of Zintl phases can be tuned to optimize thermoelectric performance.¹⁸

3 Zintl phase thermoelectrics

An example of the flexibility of a Zintl phase with potentially good thermoelectric properties is the solid solution $\text{Ca}_x\text{Yb}_{1-x}\text{Zn}_2\text{Sb}_2$ (Fig. 3). These materials contain electron precise Zn–Sb anionic

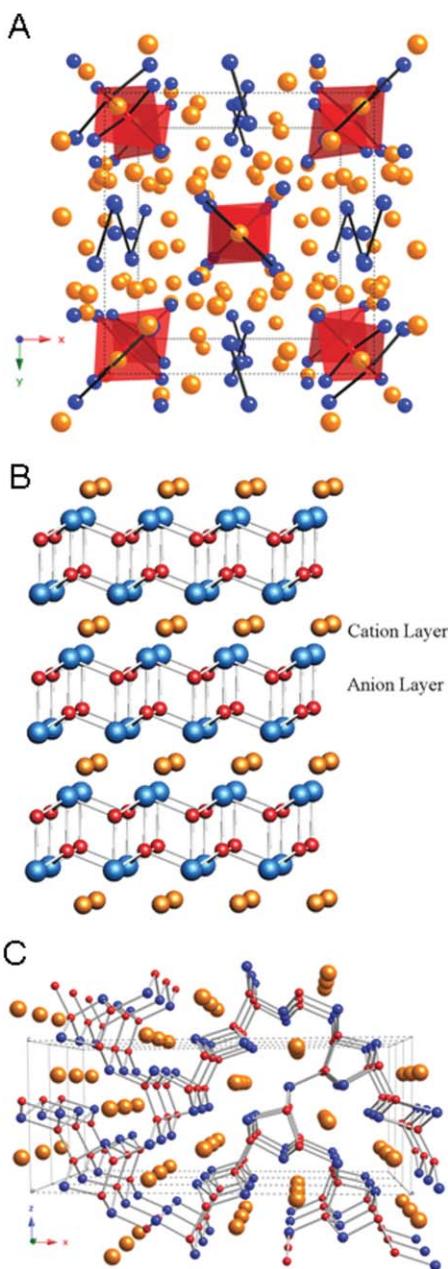


Fig. 2 Example structures of Zintl phases showing the covalently bonding regions (blue and red atoms) separate from the ionic regions (yellow atoms). (A) $\text{Ca}_{11}\text{AlSb}_9$, (B) CaZn_2Sb_2 , and (C) BaGa_2Sb_2 . The covalent bonded regions provide ‘electron–crystal’ properties such as high mobility, while the ionic region can be doped or alloyed for tuning of the electron concentration and creating disorder to give ‘phonon–glass’ properties. The gold, blue and red spheres represent the Zintl cation, antimony, and transition metal atoms, respectively.

sheets. The bonding within the sheets should have mostly covalent character due to the similar electronegativity of Zn and Sb. Formally, the structure consists of A^{2+} cations which donate electrons to the $(\text{Zn}_2\text{Sb}_2)^{2-}$ framework. The $(\text{Zn}_2\text{Sb}_2)^{2-}$ framework has good hole mobility due to its low polarity, in contrast to the polar covalent Mg_3Sb_2 with similar structure which despite some promising reports, is a poor thermoelectric.²⁰ The bonding to the Ca site is ionic and CaZn_2Sb_2 is a semiconductor with a band

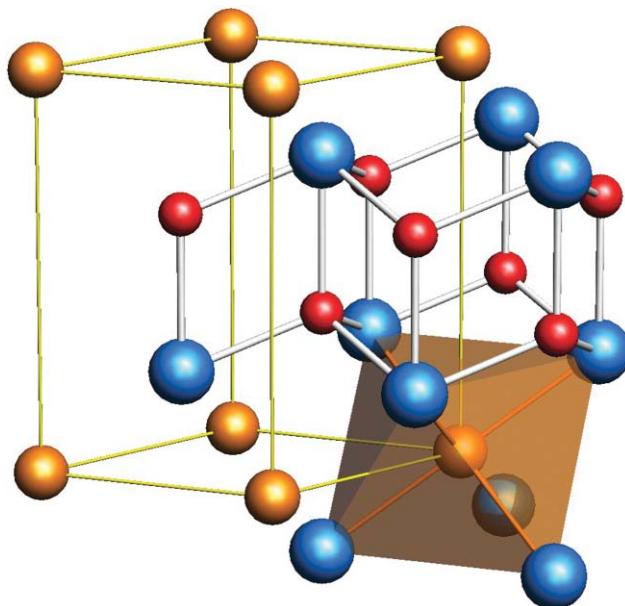


Fig. 3 A view of the structure of CaZn_2Sb_2 showing the bonding within the $\text{Zn}_2\text{Sb}_2^{2-}$ layers between planes of Ca^{2+} . The covalent bonding in the $\text{Zn}_2\text{Sb}_2^{2-}$ layers provides ‘electron–crystal’ properties such as high mobility, while substitutions (with Yb for instance) can be made at the Ca site for tuning of the electron concentration and creating disorder for ‘phonon–glass’ properties. The gold, blue and red spheres represent Ca, Sb, and Zn atoms, respectively.

gap of about 0.25 eV. The less electropositive Yb cation makes YbZn_2Sb_2 slightly electron deficient so that it is a heavily doped p-type semiconductor with metal-like transport properties. The complete solid solution $\text{Ca}_x\text{Yb}_{1-x}\text{Zn}_2\text{Sb}_2$ forms and the material transitions from a classical Zintl phase (Ca-rich semiconductor) to a heavily doped polar intermetallic (Yb-rich metal). Aside from carrier concentration, the electronic band structure of the conducting Zn_2Sb_2 framework does not change appreciably as evidenced by the invariance of the hole effective mass as a function of x in $\text{Ca}_x\text{Yb}_{1-x}\text{Zn}_2\text{Sb}_2$.¹⁸ Thus, by alloying *isoelectronic* species which different donor characteristics on the cationic site, a gradual tuning of the hole concentration is enabled which permits the optimization of zT . In addition, the disorder scattering introduced by alloying Ca and Yb on the cation site noticeably reduces the lattice component of the thermal conductivity. However, the relatively simple structure prevents the thermal conductivity from being low enough to achieve $zT > 1$.

Zn_4Sb_3 is another example of a high zT material recently identified as a complex, electron precise ($\text{Zn}_{13}\text{Sb}_{10}$) structure²¹ with two types of antimony atoms: Sb–Sb bonded dumbbells, Sb_2^{4-} , and isolated Sb^{3-} without bonds to other antimony atoms. In the Zintl formulation, it would be considered to contain 13 Zn^{2+} cations along with 6 Sb^{3-} anions, and 2 Sb_2^{4-} dumbbells. However, this ionic description is not reflected in the true electronic structure, which is better described by complex, delocalized multicenter bonding,²² which imposes a local structure visible in the low temperature polymorph. The room temperature Zn_4Sb_3 phase maintains this delocalized multicenter bonding in the local structure²³ but at longer length scales (>10 Å) the Zn atoms appear disordered. Thus, in the high-symmetry room temperature structure, 20% of the Zn atoms appear in interstitial sites (Fig. 4). The high zT

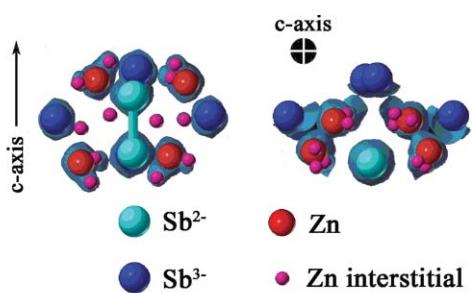


Fig. 4 A view of the structure of Zn_4Sb_3 including an electron density map that clearly shows the existence of interstitial Zn atoms.

in Zn_4Sb_3 can be traced to the low lattice thermal conductivity (less than $1 \text{ W m}^{-1} \text{ K}^{-1}$) which is attributed to the complex, Zn-interstitial containing structure.

Binary Zintl compounds containing Group 1 or 2 cations can also form complex structures providing phonon–glass qualities. A good example is the $\text{A}_{11}\text{Sb}_{10}$ structure type^{34–36} which can be described by the Zintl formulism as containing 44A²⁺ cations with two $[\text{Sb}_4]^{4-}$ squares, eight $[\text{Sb}_2]^{4-}$ dumbbells, and 16 Sb³⁻ anions (Fig. 5) in each unit cell. Indeed this complex structure with multiple anion types has remarkably low lattice thermal conductivity (about $1 \text{ W m}^{-1} \text{ K}^{-1}$) for such a high melting point compound (1200 K).²⁴ Because the electrons are well accounted for in the bonding, the electronic structure is like that of a semiconductor satisfying the electron–crystal requirement. $\text{Ca}_{11}\text{Sb}_{10}$ behaves like a lightly doped semiconductor while $\text{Yb}_{11}\text{Sb}_{10}$ is more metallic like a heavily doped semiconductor. Similar to

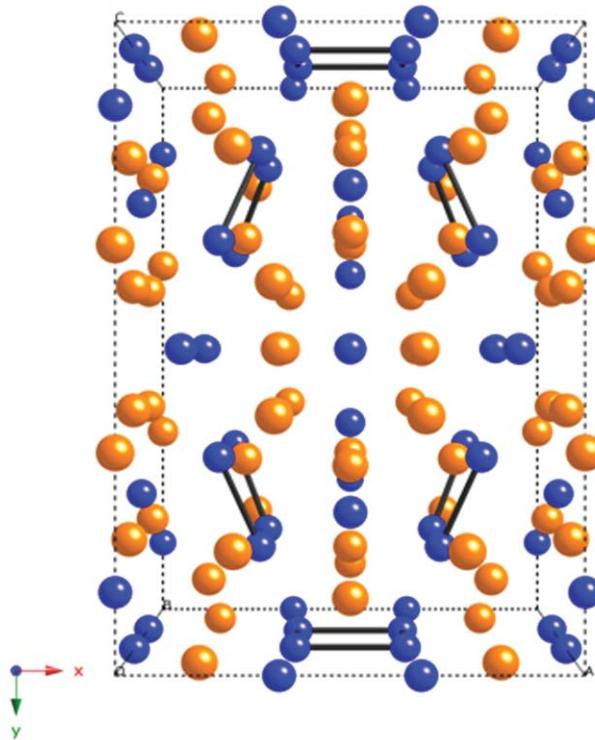


Fig. 5 A view of the structure of $\text{A}_{11}\text{Sb}_{10}$ showing the 3 distinct types of Sb atoms: Sb square rings, Sb dumbbells and isolated Sb ions. A and Sb are designated by gold and blue spheres.

$\text{Yb}_x\text{Ca}_{1-x}\text{Zn}_2\text{Sb}_2$, precise tuning of the semiconductor to metal electronic properties should be possible in $\text{Yb}_x\text{Ca}_{11-x}\text{Sb}_{10}$. The $\text{A}_{11}\text{Sb}_{10}$ materials, however, show signs of bipolar conduction (containing both free electrons and holes) which leads to low, compensated thermopower and therefore low zT .

Indeed, a number of materials identified as good thermoelectrics contain structures and bonding much more complicated than the simple alloys of traditional thermoelectric materials, and therefore benefit from a Zintl or polar intermetallic description of their structures, such as clathrates,⁷ which are not described in this short perspective. Similar to the example provided above with CaZn_2Sb_2 , the electron transport properties can be finely tuned and optimized for thermoelectric applications.

4 Transition metal Zintl phases

Antimonide skutterudites,²⁵ for example, have been found to be good thermoelectric materials with maximum zT in the 0.8 to 1.1 range for both n-type and p-type. The skutterudites structure contains a square ring of Sb atoms at typical Sb–Sb bonding distances so that this unit can be considered a Sb_4^{4-} Zintl anion with formal oxidation state of Sb^{1-} . Because of the similar electronegativity with the transition metal (commonly Co or isoelectronic species) the interaction between the antimony and the transition metal will be strongly covalent, producing a high mobility compound semiconductor. The structure allows a number of alloying possibilities on both the transition metal and Sb sites. In addition, there is a rather large empty void in the skutterudite that will readily accept a large alkaline earth or rare earth cation, making what is known as a filled skutterudite (shown in Fig. 6). However, solubility of alloys and filling atoms is limited to a narrow range of compositions nearly isoelectronic to CoSb_3 (within about $1/4\text{e}^-$ per transition metal). Thus, the materials

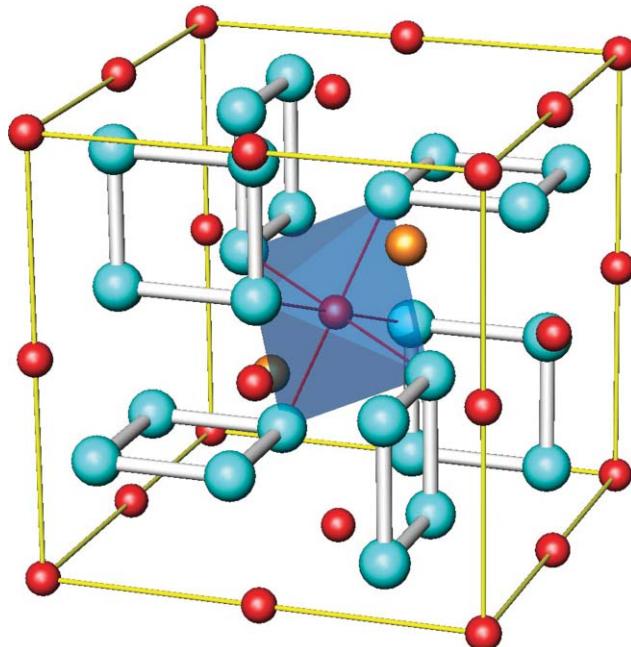


Fig. 6 A view of the structure of the filled skutterudite emphasizing the Sb_4^{4-} square ring and the metal centered octahedron. The gold, blue, and red spheres represent the cation, antimonide, and metal atoms, respectively.

are nearly electron precise qualifying the filled compositions as transition metal Zintl phases. The disorder and ‘rattling’ of the filling atom has been associated with a reduction in lattice thermal conductivity and is another example (in addition to AZn_2Sb_2 described above) of the ability of Zintl phases to separate electron–crystal regions from phonon–glass regions.

Materials crystallizing in the Ir_3Ge_7 structure, such as Mo_3Sb_7 , (shown in Fig. 7), have also been shown to be promising for thermoelectric applications. Mo_3Sb_7 , like the antimonide skutterudites, is capable of achieving a low lattice thermal conductivity and good Seebeck coefficient, resulting in a good figure of merit. Mo_3Sb_7 is a complex three-dimensional arrangement of atoms comprised of antimony dimers, square antiprisms and empty Sb_8 cubes.^{26,27} The Sb_8 cubes have the potential to house interstitial elements, opening opportunities for the addition of dopants. In addition, the structure of Mo_3Sb_7 is reported to be able to accommodate a number of different valence-electron-counts (51 to 56) and elemental constituents possibly further aiding in the optimization of its thermoelectric properties. If one counts only the short Sb–Sb dimer distance as a bond (these bonds are emphasized in Fig. 7), the valence of Mo_3Sb_7 can be rationalized as being two electrons short of valence balance ($\text{Mo}_3^{5+}\text{Sb}_4^{2-}\text{Sb}_3^{3-}$). Substitution of Te for Sb will add an electron and bring the system closer to valence balance. This accounts for the p-type, nearly semiconducting behavior of $\text{Mo}_3\text{Sb}_{7-x}\text{Te}_x$ as x approaches 2. For example, the addition of Te ($x = 1.6$) into $\text{Mo}_3\text{Sb}_{7-x}\text{Te}_x$ increases the Seebeck coefficient and decrease the thermal conductivity. The combination of resistivity, Seebeck coefficient and thermal conductivity results in a maximum zT of ~ 0.8 at 1050 K.²⁶ Much efforts are in progress to tune this material further, with attention being focused on modifying the carrier concentration (a.k.a. controlling of the valence-electron-count) and filling the Sb_8 cube voids. Mo_3Sb_7 shows yet another example of a complex transition metal Zintl phase capable of achieving an extremely reduced lattice thermal conductivity and good thermoelectric properties. However, even though the reduced lattice thermal conductivity of Mo_3Sb_7 is good for thermoelectric applications,

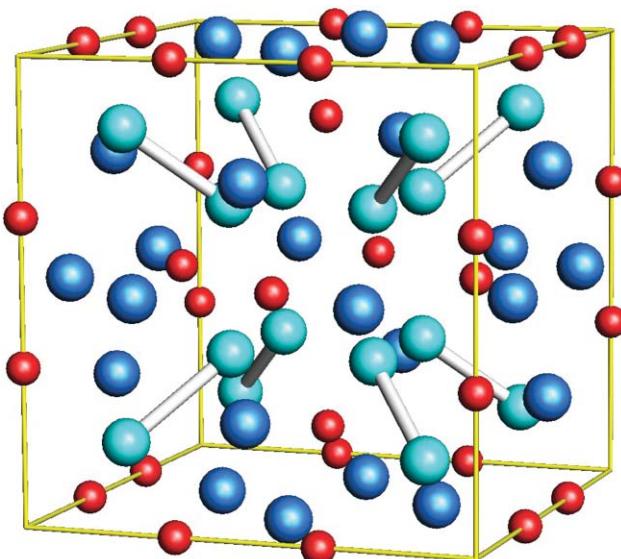


Fig. 7 A view of the Mo_3Sb_7 structure where the Sb–Sb bonds (light blue) and isolated Sb (dark blue) and Mo (red) are shown.

the recently published thermoelectric data for $\text{Yb}_{14}\text{MnSb}_{11}$ reveals an even lower lattice thermal conductivity.

The recent discovery of high zT in $\text{Yb}_{14}\text{MnSb}_{11}$ demonstrates the successful combination of electron–crystal phonon–glass properties afforded by Zintl phases.⁸ $\text{Yb}_{14}\text{MnSb}_{11}$ is composed of distinct anionic units (like $\text{Yb}_{11}\text{Sb}_{10}$) rather than the extended structures shown in Fig. 2 (such as CaZn_2Sb_2). The structure consists of (Fig. 8) $[\text{MnSb}_4]^{9-}$ tetrahedra, polyatomic $[\text{Sb}_3]^{7-}$ anions, as well as isolated Sb^{3-} anions, where the electrons are donated by the Yb, to give Yb^{2+} cations. This compound is isostructural to the well-known Zintl phase, $\text{Ca}_{14}\text{AlSb}_{11}$,²⁸ with ionic bonds between the cation and the anionic units. In principle, since these anionic units are ionically bonded together through cations, a high temperature melting point is also expected, providing incentive for high temperature stable thermoelectric properties. While this structure is clearly much more complex than the example provided above for CaZn_2Sb_2 , it has many of the same attributes and provides for the possibilities of fine tuning.

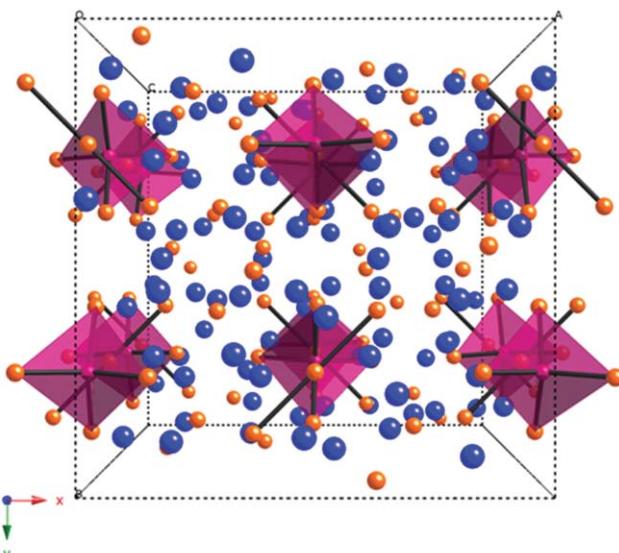


Fig. 8 A view of $\text{A}_{14}\text{MPn}_{11}$ structure ($\text{A} = 2+$ cation, $\text{M} = \text{metal}$, $\text{Pn} = \text{P}$, As, Sb, Bi), where the MPn_4 polyhedra and Pn_3 unit are shown. The gold, blue, and red spheres represent the A , Pn , and M atoms, respectively.

The generic expression for this structure type is $\text{A}_{14}\text{MPn}_{11}$, where A is a heavy or alkaline earth metal, M is a transition or main group metal, and Pn is a heavier Group 15 element (P, As, Sb, or Bi).^{10,16} The electronic structure of the main group analog is consistent with the Zintl formulism.²⁹ The various sites allow for the potential tuning of the electronic and thermal properties through elemental substitutions. Substituting on the cationic metal site, A , allows for possible tuning of the carrier concentration and disorder scattering of phonons, while substituting on the metal site, M , enables possible tuning of the electronic parameters.

$\text{Yb}_{14}\text{MnSb}_{11}$ is best considered as a polar intermetallic (because of the metallic transport properties)^{30,31} compared with the transition metal Zintl phase, $\text{Ca}_{14}\text{MnSb}_{11}$ (semiconducting).³² However, the presence of the transition metal changes the electronic structure and bonding within this structure type.³³ X-Ray magnetic circular dichroism (XMCD)³⁴ and XPS³⁵ measurements show that, indeed, the ytterbium is in the Yb^{2+} valence state replacing Ca^{2+} ;

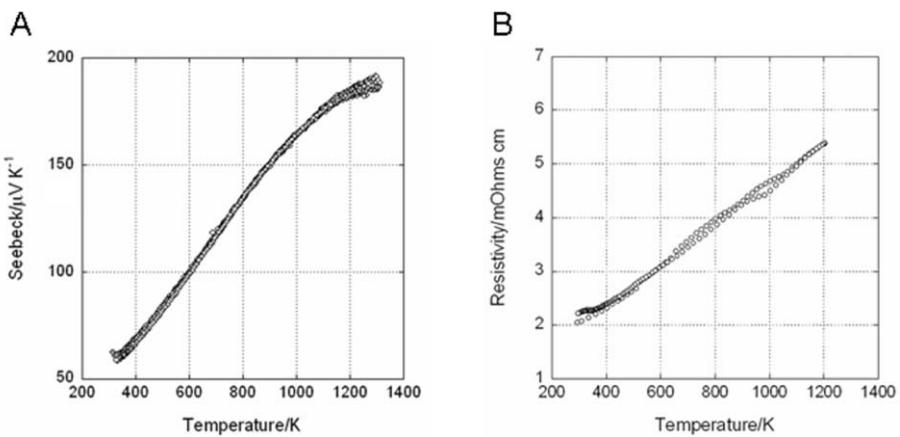


Fig. 9 (A) Seebeck and (B) resistivity vs. temperature for $\text{Yb}_{14}\text{MnSb}_{11}$.

however, the manganese has been shown to be in the Mn^{2+} valence state, supplying one less electron than Al^{3+} . This shortage of electrons ($1.3 \times 10^{21} \text{ holes cm}^{-3}$) explains the material's p-type behavior and corresponds well to the measured carrier concentration by Hall effect. This high carrier concentration indicates that the hole mobility is not particularly high, only about $3 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$, but the high thermopower implies a large effective mass $m^* \sim 3 m_e$ using eqn (1). Fig. 9 shows the temperature dependent Seebeck coefficient and resistivity. The Seebeck coefficient is positive and increases with increasing temperature. The resistivity increases with increasing temperature and reaches a maximum at about $5 \text{ m}\Omega \text{ cm}$ at 1200 K. These measurements also show that this is a stable high temperature material, with no melting or structural transition up to 1200 K. A reduction in carrier concentration, by substitution of La^{3+} for Yb^{2+} on the A site or by replacing Mn^{2+} with Al^{3+} should bring the peak in zT to lower temperatures, providing higher average zT over a wider temperature range. In addition, isoelectronic alloying on the A site, for example by replacing some Yb^{2+} with Ca^{2+} should result in a further reduction in lattice thermal conductivity as seen in $\text{Yb}_{1-x}\text{Ca}_x\text{Zn}_2\text{Sb}_2$.¹⁸

The total thermal conductivity of $\text{Yb}_{14}\text{MnSb}_{11}$, shown in Fig. 10, is remarkably low ranging between ~ 0.7 – $0.9 \text{ W m}^{-1} \text{ K}^{-1}$ for the temperatures of 300–1275 K. Subtracting the electronic contribution results in an extremely low lattice thermal conductivity. This

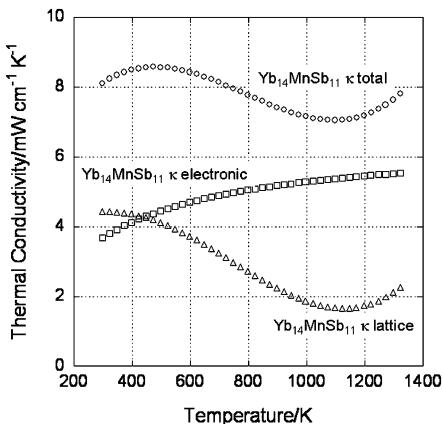


Fig. 10 Thermal conductivity vs. temperature for $\text{Yb}_{14}\text{MnSb}_{11}$. Total thermal conductivity, electronic thermal conductivity, and lattice thermal conductivity are shown.

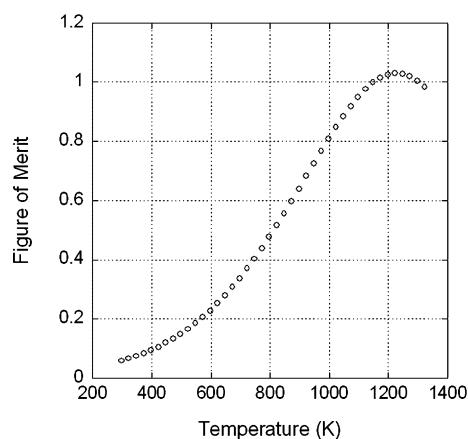


Fig. 11 zT , Figure of merit, vs. temperature for $\text{Yb}_{14}\text{MnSb}_{11}$.

thermal conductivity is comparable to that of a glass,³⁶ heavily contributing to the material's high figure of merit (Fig. 11). The low values can be attributed to the complexity of the structure (limiting the phonon mean-free path) and heavy atomic mass (reducing the fraction of atomic vibrational modes that carry heat efficiently) of the crystal.

$\text{Yb}_{14}\text{MnSb}_{11}$ possesses a high figure of merit ($zT \geq 1$) in the temperature region (up to 1275 K) where the state-of-the-art material is p-Si_{1-x}Ge_x alloy (SiGe), with a maximum $zT = 0.6$.^{37,38} Fig. 12 compares the most competitive p-type materials, showing the improvement of $\text{Yb}_{14}\text{MnSb}_{11}$ in the high temperature range. For highest efficiency $\text{Yb}_{14}\text{MnSb}_{11}$ would be most effectively utilized in a segmented thermoelectric device where other materials with high zT at lower temperatures such as TAGS are used in the lower temperature region.³⁹ In order to utilize a segmented element, a thermoelectric compatibility factor, s , must also be considered. The compatibility factor takes into account the requirement that the same electrical current and similar heat must flow through each segment. Fig. 13 shows the compatibility factor, s , for materials being considered for such a device. If the compatibility factors differ by more than a factor of 2, efficiency of the device is substantially reduced. $\text{Yb}_{14}\text{MnSb}_{11}$ is much more compatible than the current high temperature p-type material, SiGe, for segmentation with lower temperature high zT materials such as TAGS because of the closer compatibility factor.

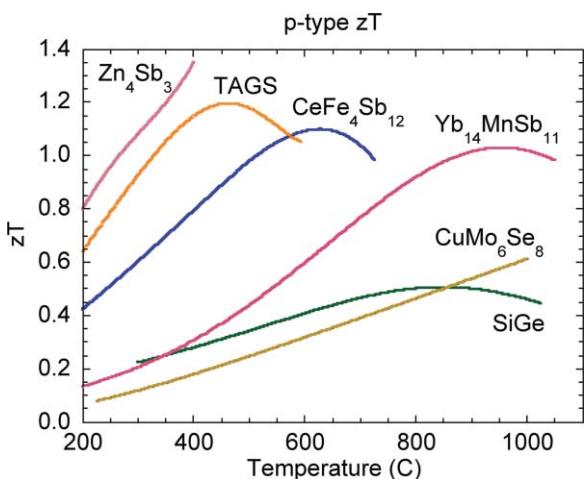


Fig. 12 Figure of merit as a function of temperature for a number of p-type materials (TAGS: $(\text{GeTe})_{0.85}(\text{AgSbTe}_2)_{0.15}$). Reprinted with permission from *Chem. Mater.*, 2006, **18**, 1873. © American Chemical Society 2006.

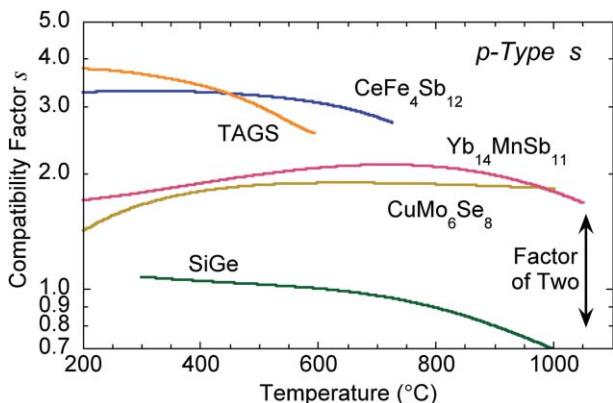


Fig. 13 Thermoelectric compatibility factor, s , for a variety of p-type materials. $\text{Yb}_{14}\text{MnSb}_{11}$ compatibility is improved by a factor of 2 over the current material, SiGe. (TAGS: $(\text{GeTe})_{0.85}(\text{AgSbTe}_2)_{0.15}$). Reprinted with permission from *Chem. Mater.*, 2006, **18**, 1873. © American Chemical Society 2006.

In addition, because of the many known compositions with the same structure type, it also introduces a chemically rich, bulk material into the thermoelectric community that can be further optimized.

5 Conclusions and outlook

Zintl phases possess the structural characteristics needed for electron–crystal phonon–glass properties that lead to high thermoelectric efficiency (zT). They combine distinct regions of covalent bonding ideal for electron–crystal properties with ionically bonded cations that can be easily substituted for precise tuning of electronic properties and introducing disorder. The complex structures and disorder creates phonon–glass properties resulting in low lattice thermal conductivity, essential for high zT . A good understanding of the interaction of the bonding and chemistry with the transport properties in these compounds leads to a focused search for new materials and a rational approach to optimizing the electronic and thermal properties.

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