

Inorganic Chemistry

including bioinorganic chemistry

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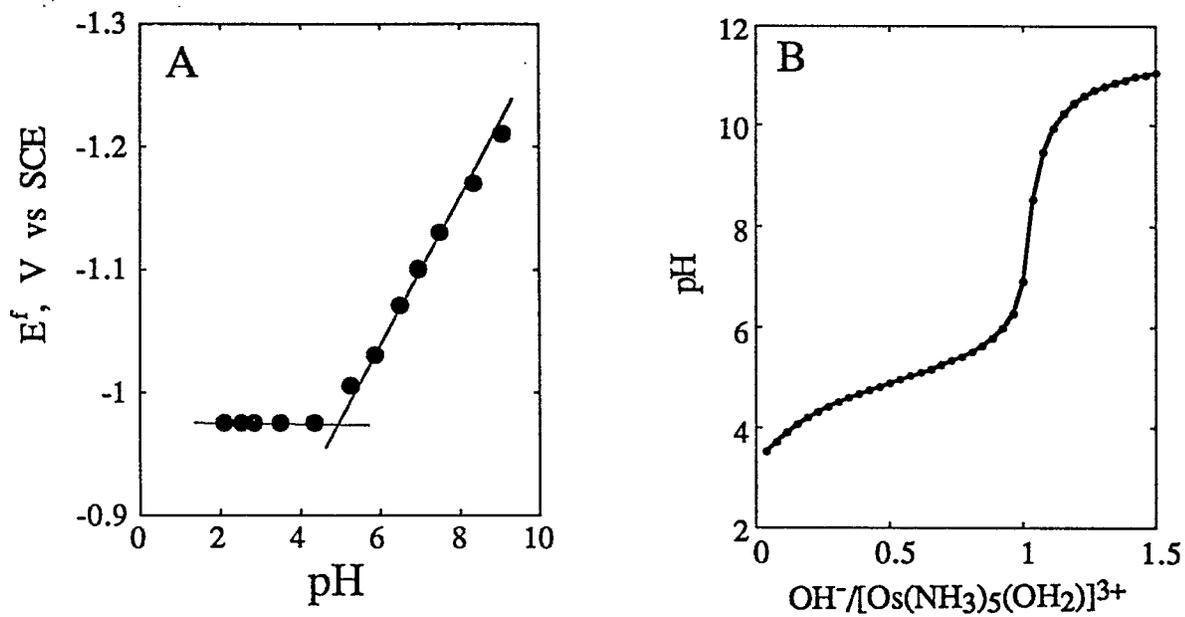


Figure S1. (A) Variation of the formal potential of the $[\text{Os}(\text{NH}_3)_5(\text{OH}_2)]^{3+/2+}$ couple with the pH of the supporting electrolytes which were buffered between pH 2.1 and 9.1. (B) pH titration curve for the titration of 5.4 mM $[\text{Os}(\text{NH}_3)_5(\text{OH}_2)]^{3+}$ with 0.01 M NaOH.

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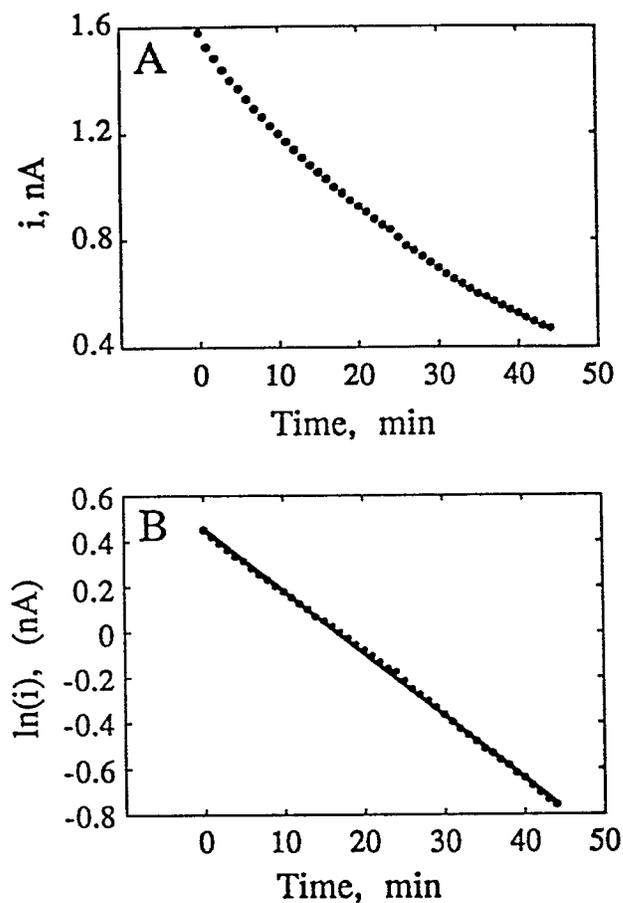


Figure S2. (A) Time dependence of the steady-state current for the oxidation of $[\text{Os}(\text{NH}_3)_5(\text{OH}_2)]^{2+}$ at a carbon fiber microelectrode maintained at -0.7 V. The initial concentration of Os(II) was 1.3 mM. Supporting electrolyte: 0.1 M $\text{CH}_3\text{SO}_3\text{Na}$. Initial pH = 5.6 . (B) First-order plot of the data from (A).

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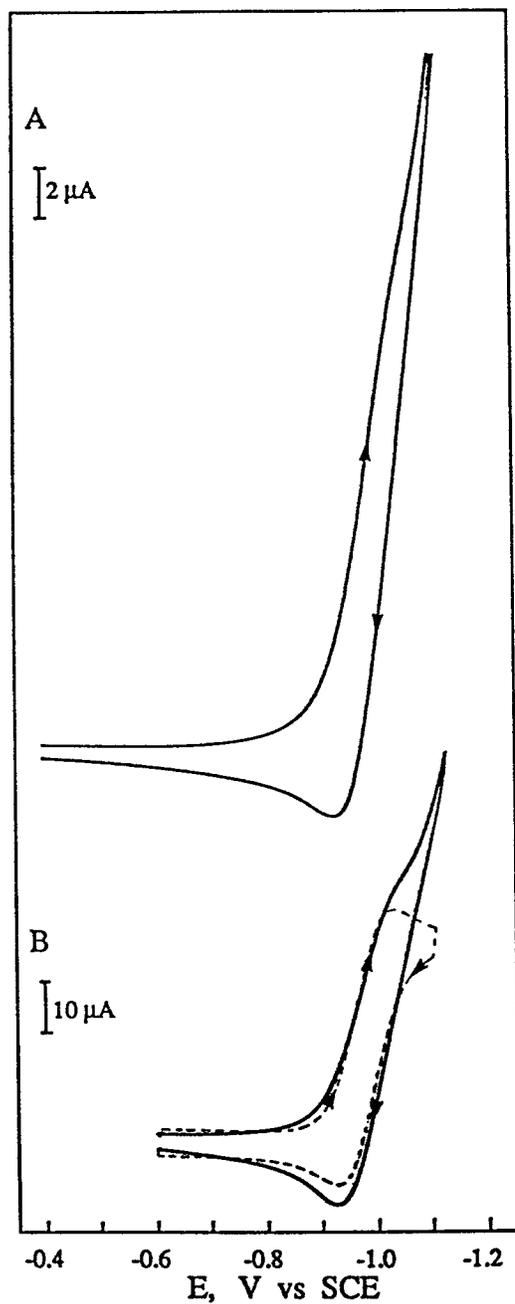


Figure S3. (A) Cyclic voltammogram for the reduction of 0.74 mM $[\text{Os}(\text{NH}_3)_5(\text{OH}_2)]^{3+}$ in $0.1 \text{ M CH}_3\text{SO}_3\text{Na} + 0.48 \text{ M CH}_3\text{COOH}$ ($\text{pH} = 2.53$). Scan rate: 5 mV s^{-1} . (B) Repeat of (A) at a scan rate of 50 mV s^{-1} . The dashed curve was calculated with the Digisim program (see text).